

pH Measurements – Buffers and Their Properties

1. A solution of a weak acid was tested with the indicators used in this experiment. The colors observed were as follows:

Methyl violet	<i>violet</i>
Thymol blue	<i>yellow</i>
Methyl yellow	<i>yellow</i>
Congo red	<i>orange-red</i>
Bromocresol green	<i>green</i>

What is the approximate pH of the solution? _____

2. A solution of Na_2CO_3 has a pH of 10. The CO_3^{2-} ion is the conjugate base of the HCO_3^- ion. Write the net ionic equation for the reaction which makes a solution of Na_2CO_3 basic. (See Equation 7.11)

3. The pH of a 0.10 M HCN solution is 5.20.

a. What is $[\text{H}_3\text{O}^+]$ in that solution?

_____ M

b. What is $[\text{CN}^-]$? What is $[\text{HCN}]$? (Where do the H^+ and CN^- ions come from?)

$[\text{CN}^-] =$ _____ M; $[\text{HCN}] =$ _____ M

c. What is the value of K_a for HCN? What is the value of $\text{p}K_a$?

$K_a =$ _____ $\text{p}K_a =$ _____

Ph Measurements Buffers And Their Properties

Hussin A.Rothana



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