

ACTIVITY: Molecular Polarity Computer Simulation (PhET)

Go to: <https://openstax.org/r/phet> Click on the arrow, DOWNLOAD, OPEN and RUN the MOLECULE POLARITY program. Or google "phet molecule polarity" to find the simulation.

Introduction:

In this atomic level simulation, you will investigate how atoms' **electronegativity** values affect the bonds they produce. When two atoms form a covalent bond, a pair of electrons is shared between atoms. **Electronegativity** is a measure of a single atom's ability to attract electrons shared in that bond. In this lab you will work at your own pace, to answer a number of questions. To begin, from what you've already learned about the protons and electrons in an atom, what would cause an atom to have a **high** electronegativity value?

Why might an atom have a low electronegativity value?

Procedure:

New Molecule

- ☛ Turn on (check) all view options. **distort** our gaze the molecule with your mouse.
- ☛ Take your time and investigate how the binary compound's bond behaves when the atom's electronegativity and orientation are changed. **Do not rush through this step.** Describe the bond formed between two atoms with **similar, low** electronegativities.

Describe the bond formed between two atoms with **similar, high** electronegativities.

Describe the bond formed between two atoms with **very different** electronegativities.

Describe (in your own words) what is meant by **partial charges**, δ^- and δ^+ .

δ^- represents: _____

δ^+ represents: _____

What happens when the electric field is applied to a very polar molecule?

Why do you think this is? _____

View

- ☒ Bond Dipoles **ON**
- ☒ Partial Charges
- ☒ Bond Orientation

Surfaces

- ☐ None
- ☐ Electrostatic Potential
- ☒ Electron Density

Electric Field

- ☒ ON ☐ OFF



Phet Molecular Polarity Lab Answer Key

Nivaldo J. Tro



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New to MasteringChemistry: MasteringChemistry metadata analysis of problems/tutorials assigned in the previous edition have been used to revise end-of-chapter problems in the Third Edition. Approximately 1,000 end-of-chapter questions have been enhanced with feedback meeting instructor's need for more tutorial-like questions. Interactive versions of selected worked examples in the text have been created and are incorporated into MasteringChemistry as assignable tutorial activities, providing an office-hour-like experience. These can also be used for mobile learning through a downloadable app.

15. Pause and Predict Video Quizzes bring chemistry to life with lab demonstrations illustrating key topics in general chemistry. Students are asked to predict the outcome of experiments as they watch the videos; a set of multiple-choice questions challenges students to apply the concepts from the video to related scenarios.

8. PhET tutorials have been developed around interactive applets that foster conceptual understanding and active learning. Topics include acid-base solutions, balancing chemical equations, and molecular polarity. Multiple-choice Reading Questions are provided for each chapter, making it easy to hold students accountable for doing assigned readings before lecture. Enhanced end-of-chapter questions within MasteringChemistry, providing wrong-answer feedback, have been added. Sketch-it-type problems have been added for each chapter. Simulations cover some of the most difficult chemistry concepts and are written by the leading authors in simulation development. Select end-of-chapter questions and reading quizzes have been tagged to learning outcomes. The overall number of algorithmic and randomized problems has been increased to 40%, offering a more rounded program for departments moving to online high-stakes testing.

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Table of Contents Phet Molecular Polarity Lab Answer Key

1. Understanding the eBook Phet Molecular Polarity Lab Answer Key
 - The Rise of Digital Reading Phet Molecular Polarity Lab Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Phet Molecular Polarity Lab Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Phet Molecular Polarity Lab Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Phet Molecular Polarity Lab Answer Key

- Personalized Recommendations
 - Phet Molecular Polarity Lab Answer Key User Reviews and Ratings
 - Phet Molecular Polarity Lab Answer Key and Bestseller Lists
5. Accessing Phet Molecular Polarity Lab Answer Key Free and Paid eBooks
 - Phet Molecular Polarity Lab Answer Key Public Domain eBooks
 - Phet Molecular Polarity Lab Answer Key eBook Subscription Services
 - Phet Molecular Polarity Lab Answer Key Budget-Friendly Options
 6. Navigating Phet Molecular Polarity Lab Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Phet Molecular Polarity Lab Answer Key Compatibility with Devices
 - Phet Molecular Polarity Lab Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Phet Molecular Polarity Lab Answer Key
 - Highlighting and Note-Taking Phet Molecular Polarity Lab Answer Key
 - Interactive Elements Phet Molecular Polarity Lab Answer Key
 8. Staying Engaged with Phet Molecular Polarity Lab Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Phet Molecular Polarity Lab Answer Key
 9. Balancing eBooks and Physical Books Phet Molecular Polarity Lab Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Phet Molecular Polarity Lab Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Phet Molecular Polarity Lab Answer Key
 - Setting Reading Goals Phet Molecular Polarity Lab Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Phet Molecular Polarity Lab Answer Key

- Fact-Checking eBook Content of Phet Molecular Polarity Lab Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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