

# Key

## Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant  $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$  to solve the following problems:

If pressure is needed in kPa then convert by multiplying by  $101.3 kPa / 1 atm$  to get  
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87 °C, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56 °C, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of 67 °C, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

# Phet Idael Gas Law Answer Key

**Patrick Hurley**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Phet Idael Gas Law Answer Key:**

*A Guide to Teaching in the Active Learning Classroom* Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment, for administrators planning to create ALCs or experimenting with provisionally designed rooms, and for faculty developers helping teachers transition to using these new spaces.

**Simulations and Student Learning** Matthew Schnurr, Anna MacLeod, 2021-01-04 The book underlines the value of simulation-based education as an approach that fosters authentic engagement and deep learning.

*The Ideal Gas Law Handbook - Everything You Need to Know about Ideal Gas Law* Patrick Hurley, 2016-04-29 This book is your ultimate Ideal gas law resource. Here you will find the most up-to-date information, facts, quotes, and much more. In easy-to-read chapters with extensive references and links to get you to know all there is to know about Ideal gas law's whole picture right away. Get countless Ideal gas law facts right at your fingertips with this essential resource. The Ideal gas law Handbook is the single and largest Ideal gas law reference book. This compendium of information is the authoritative source for all your entertainment, reference, and learning needs. It will be your go-to source for any Ideal gas law questions. A mind-tickling encyclopedia on Ideal gas law, a treat in its entirety, and an oasis of learning about what you don't yet know but are glad you found. The Ideal gas law Handbook will answer all of your needs and much more.

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### Table of Contents **Phet Idael Gas Law Answer Key**

1. Understanding the eBook **Phet Idael Gas Law Answer Key**
  - The Rise of Digital Reading **Phet Idael Gas Law Answer Key**
  - Advantages of eBooks Over Traditional Books
2. Identifying **Phet Idael Gas Law Answer Key**
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an **Phet Idael Gas Law Answer Key**
  - User-Friendly Interface
4. Exploring eBook Recommendations from **Phet Idael Gas Law Answer Key**
  - Personalized Recommendations
  - **Phet Idael Gas Law Answer Key** User Reviews and Ratings
  - **Phet Idael Gas Law Answer Key** and Bestseller Lists

5. Accessing Phet Idael Gas Law Answer Key Free and Paid eBooks
  - Phet Idael Gas Law Answer Key Public Domain eBooks
  - Phet Idael Gas Law Answer Key eBook Subscription Services
  - Phet Idael Gas Law Answer Key Budget-Friendly Options
6. Navigating Phet Idael Gas Law Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Phet Idael Gas Law Answer Key Compatibility with Devices
  - Phet Idael Gas Law Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Phet Idael Gas Law Answer Key
  - Highlighting and Note-Taking Phet Idael Gas Law Answer Key
  - Interactive Elements Phet Idael Gas Law Answer Key
8. Staying Engaged with Phet Idael Gas Law Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Phet Idael Gas Law Answer Key
9. Balancing eBooks and Physical Books Phet Idael Gas Law Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Phet Idael Gas Law Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Phet Idael Gas Law Answer Key
  - Setting Reading Goals Phet Idael Gas Law Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Phet Idael Gas Law Answer Key
  - Fact-Checking eBook Content of Phet Idael Gas Law Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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