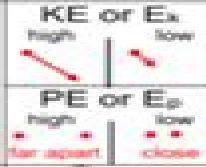


Study Buddy: Energy and Phase Changes

Terms	Description	Picture	Symbol/Unit
Heat	Amount of thermal energy transferred from one system to another. http://openstax.org/r/media_img/video/FILE:Heat1_1st.png http://www.ck12.org/ck12-quartz/277501/277501-048-001.png		Q/ Joules (J) or calorie (cal)
Total Internal Energy	Describes the total amount of energy in a system. Equals the sum of kinetic and potential energies. Kinetic Energy: due to motion of particles Potential Energy: due to relative position of particles		IE/ Joules (J) or calorie (cal)
Endothermic Process	System absorbs energy from surroundings Phase changes: melting, vaporizing, subliming http://openstax.org/r/ck12-quartz/277501/277501-048-001.png		+Q for system
Exothermic Process	System releases heat to surroundings Phase changes: freezing, condensing, depositing		-Q for system
Vapor Pressure	Pressure exerted by the vapor particles over a liquid or solid substance at equilibrium.		http://www.ck12.org/ck12-quartz/277501/277501-048-001.png

How does thermal energy transfer cause phase changes?

As a solid or liquid is heated, the thermal energy is used to increase the kinetic and potential energy of the particles (particles moving faster and farther apart). When a liquid or gas is cooled, the thermal energy transferred to the surroundings reduced the kinetic and potential energies of the particles.



Describe solid particles' positions and motion:

- Moving slowly
- Close together

Solid water



Describe liquid particles' positions and motion:

- Moving faster
- Farther apart

Liquid water



Describe gaseous particles' positions and motion:

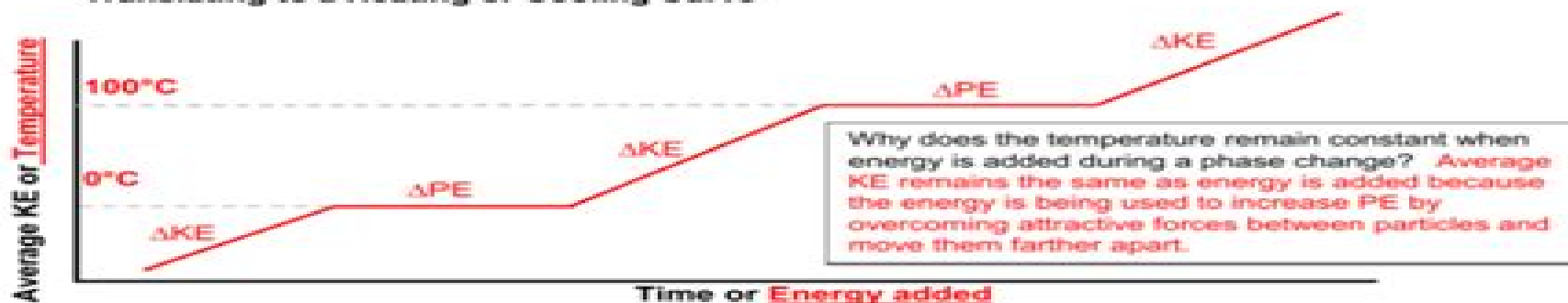
- Moving quickly
- Far apart

Water vapor

Melting = 0 °C

Boiling = 100 °C

Translating to a Heating or Cooling Curve



Label where PE and KE are changing on the heating curve above.

Phase Changes Study Guide

DJ Losen



Phase Changes Study Guide:

Physics for Scientists and Engineers Study Guide Todd Ruskell, 2007-06-08 The Study Guide provides students with key physical quantities and equations misconceptions to avoid questions and practice problems to gain further understanding of physics concepts and quizzes to test student knowledge of chapters

Chemistry, Student Study Guide James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Chemical Principles Student's Study Guide & Solutions Manual John Krenos, Joseph Potenza, Loretta Jones, Lynn Koplitz, Thomas Spence, 2004-03-19 This combination manual is designed to help students avoid common mistakes and understand the material better The solutions manual section includes detailed answers and explanations to the odd numbered exercises in the text

Chemical Principles Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications

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advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **CompTIA**

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