

The letters "pH" represent the French words "pouvoir hydrogene" which means "hydrogen power".

- The definition of pH is *pH is equal to the negative log (logarithm) of the hydrogen ion concentration of a solution.*
- The logarithm of a number is the power to which 10 must be raised to equal that number.

A pH value of less than 7 indicates a(n) acidic solution. A pH value of 7 indicates a neutral solution. A pH value of more than 7 indicates a(n) basic solution.

PROBLEMS: Show all work and circle the final answer.

1. Determine the pH of a 0.010 M HNO_3 solution.

$$\begin{aligned} \text{pH} &= -\log[0.010] \\ &= \boxed{2} \end{aligned}$$

2. What is the pH of a 2.5×10^{-6} M solution of HCl?

$$\begin{aligned} \text{pH} &= -\log[2.5 \times 10^{-6}] \\ &= \boxed{5.6} \end{aligned}$$

3. Calculate the pH of a solution of 0.0025M H_2SO_4 .

$$\begin{aligned} \text{pH} &= -\log[2 \times 0.0025] \\ &= -\log[0.005] \\ &= \boxed{2.3} \end{aligned}$$

4. Calculate the pH of a 0.0010 M NaOH solution.

$$\begin{aligned} [\text{H}^+][0.0010] &= 1.0 \times 10^{-14} \\ [\text{H}^+] &= \frac{1.0 \times 10^{-14}}{0.0010} \\ \text{pH} &= -\log[1 \times 10^{-11}] \\ &= \boxed{11} \end{aligned}$$

5. What is the pH of a 0.020M $\text{Sr}(\text{OH})_2$ solution?

$$\begin{aligned} [\text{OH}^-][0.04] &= 1 \times 10^{-14} \\ [\text{OH}^-] &= \frac{1 \times 10^{-14}}{0.04} \\ \text{pH} &= -\log[\text{H}^+] \\ &= -\log 2.5 \times 10^{-9} \end{aligned}$$

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