

## Unit 6, Lesson 08: The pH of Salt Solutions, Answers

1. Complete the following chart:

| Ion                         | Parent Acid or Base      | Is the parent strong or weak? | Will this ion hydrolyze? | If the ion will hydrolyze (react with water), write the ionization reaction.<br>If not, write "no reaction" or NR.                                      | pH of ion in solution (acidic, basic or neutral) |
|-----------------------------|--------------------------|-------------------------------|--------------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------|
| $\text{NH}_4^{1+}$          | $\text{NH}_3$            | weak                          | yes                      | $\text{NH}_4^{1+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{NH}_3(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$                   | acidic                                           |
| $\text{PO}_3^{3-}$          | $\text{H}_3\text{PO}_3$  | weak                          | yes                      | $\text{PO}_3^{3-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{HPO}_3^{2-}(\text{aq}) + \text{OH}^-(\text{aq})$                      | basic                                            |
| $\text{Na}^{1+}$            | $\text{NaOH}$            | strong                        | no                       | NR                                                                                                                                                      | neutral                                          |
| $\text{N}_2\text{H}_5^{1+}$ | $\text{N}_2\text{H}_4$   | weak                          | yes                      | $\text{N}_2\text{H}_5^{1+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{N}_2\text{H}_4(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$ | acidic                                           |
| $\text{HCO}_3^{1-}$         | $\text{H}_2\text{CO}_3$  | weak                          | yes                      | $\text{HCO}_3^{1-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{H}_2\text{CO}_3(\text{aq}) + \text{OH}^-(\text{aq})$                 | basic                                            |
| $\text{SO}_4^{2-}$          | $\text{H}_2\text{SO}_4$  | strong                        | no                       | NR                                                                                                                                                      | neutral                                          |
| $\text{Ca}^{2+}$            | $\text{Ca}(\text{OH})_2$ | strong                        | no                       | NR                                                                                                                                                      | neutral                                          |
| $\text{NO}_3^{1-}$          | $\text{HNO}_3$           | strong                        | no                       | NR                                                                                                                                                      | neutral                                          |

### Notes and explanations:

- positive ions from weak bases are strong enough acids to hydrolyze. They will donate a proton to water and form the  $\text{H}_3\text{O}^+$  ion. Because this reaction forms the  $\text{H}_3\text{O}^+$  ion, it is a  $K_a$  reaction
- negative ions from weak acids are strong enough bases to hydrolyze. They will take a proton from water, leaving the  $\text{OH}^-$  ion behind. Because this reaction forms the  $\text{OH}^-$  ion, it is a  $K_b$  reaction

2. For each of the following salts:

- identify its parent acid and parent base and indicate if they are strong or weak

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**Lauren Gardner**



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