

Matric Final Exam Time Table 2025

South Africa Page 1



basic education
Department
Basic Education
REPUBLIC OF SOUTH AFRICA



2025 NATIONAL SENIOR CERTIFICATE (NSC) OCTOBER/NOVEMBER EXAMINATIONS TIMETABLE

WEEK 1	09:00	14:00
Monday 20/10	NON-EXAMINATION DAY	
Tuesday 21/10	Computer Applications Tech P1 (3hrs) Practical	Hindi, Gujarati, Tamil, Telegu, Urdu HL P1 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu FAL P1 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu SAL P1 (2hrs) Hebrew SAL P1 (2hrs) German HL, SAL P1 (2hrs)
Wednesday 22/10	Information Technology P1 (3hrs) Practical	Arabic, French, Italian, Mandarin, Modern Greek, Serbian, Spanish SAL P1 (2hrs) Latin SAL P1 (3hrs) Portuguese HL, FAL, SAL P1 (2hrs)
Thursday 23/10	English HL P3 (3hrs) English FAL P3 (2½hrs) English SAL P3 (2½hrs)	Arabic, French, Italian, Mandarin, Modern Greek, Serbian, Spanish SAL P2 (2hrs) Latin SAL P2 (2hrs) Portuguese HL P2 (2½hrs), FAL, SAL P2 (2hr) Equine Studies (3hrs)
Friday 24/10	Afrikaans HL P3 (3hrs) Afrikaans FAL P3 (2½hrs) Afrikaans SAL P3 (2½hrs)	Hindi, Gujarati, Tamil, Telegu, Urdu HL P2 (2½hrs) Hindi, Gujarati, Tamil, Telegu, Urdu FAL P2 (2hrs) Hindi, Gujarati, Tamil, Telegu, Urdu SAL P2 (2hrs) Hebrew SAL P2 (2hrs) German HL P2 (2½hrs), SAL P2 (2hrs)
WEEK 2	09:00	14:00
Monday 27/10	isiZulu, isiXhosa, siSwati, isiNdebele HL P3 (3hrs), FAL P3 (2½hrs), SAL P3 (2½hrs)	Accounting P1 (2hrs)
Tuesday 28/10	Sepedi, Sesotho, Setswana, Xitsonga, Tshivenda HL P3 (3hrs), FAL P3 (2½hrs), SAL P3 (2½hrs) South African Sign Language HL P3 (3hrs)	Civil Technology (3hrs) Sport and Exercise Science (3hrs)
Wednesday 29/10	English HL P1 (2hrs) English FAL P1 (2hrs) English SAL P1 (2hrs)	Hindi, Gujarati, Tamil, Telegu, Urdu HL P3 (2½hrs) FAL P3 (2½hrs) Portuguese, German HL P3 (2½hrs) Portuguese FAL P3 (2½hrs)
Thursday 30/10	History P1 (3hrs)	Engineering Graphics and Design P1 (3hrs)
Friday 31/10	Mathematics P1 (3hrs) Mathematical Literacy P1 (3hrs) Technical Mathematics P1 (3hrs)	Religion Studies P1 (2hrs)
WEEK 3	09:00	14:00
Monday 03/11	Mathematics P2 (3hrs) Mathematical Literacy P2 (3hrs) Technical Mathematics P2 (3hrs)	Agricultural Sciences P1 (2½hrs)
Tuesday 04/11	isiZulu, isiXhosa, siSwati, isiNdebele HL P1 (2hrs), FAL P1 (2hrs), SAL P1 (2hrs)	Accounting P2 (2hrs) Maritime Economics (3hrs)
Wednesday 05/11	Sepedi, Sesotho, Setswana, Xitsonga, Tshivenda HL P1 (2hrs), FAL P1 (2hrs), SAL P1 (2hrs) South African Sign Language HL P1 (2hrs)	Religion Studies P2 (2hrs) Mechanical Technology (3hrs)
Thursday 06/11	isiZulu, isiXhosa, siSwati, isiNdebele HL P2 (2½hrs), FAL P2 (2½hrs), SAL P2 (1½hrs)	Business Studies P1 (2hrs)
Friday 07/11	Physical Sciences (Physics) P1 (3hrs) Technical Sciences P1 (3hrs)	History P2 (3hrs)



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November 2013 Final Exam Grade 11 Time Table

Limpopo

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November 2013 Final Exam Grade 11 Time Table Limpopo:

Unveiling the Magic of Words: A Review of "**November 2013 Final Exam Grade 11 Time Table Limpopo**"

In a world defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their capability to kindle emotions, provoke contemplation, and ignite transformative change is actually awe-inspiring. Enter the realm of "**November 2013 Final Exam Grade 11 Time Table Limpopo**," a mesmerizing literary masterpiece penned with a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction especially nitration of methyl benzoate to give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical $\times 100$ what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area theory and explanation of the nitration of bromobenzene if the temperature of the reaction mixture exceeds 60 °C

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 H_2SO_4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 Å bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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in this experiment the electrophilic aromatic substitution mechanism was used to obtain 4 nitrobromobenzene and 2 nitrobromobenzene recrystallization was used to purify 4 nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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experimental spectra reaction a fluorobenzene stock fluorobenzene 1 h 1 h 19 f 19 f 13 c nmr fid for reference only sample 1 h 1 h 19 f 19 f 13 c nmr eas nitration of fluorobenzene crude product not available for submission for credit

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lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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