#### Worksheet: Acids, Bases, and Salts

Nome KEY

Section 1

#### What are Acids, Bases, and Salts?

Ba(OH)<sub>2</sub>

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

$$HNO_3(aq) \rightarrow H^*(aq) + NO_3^*(aq)$$

$$HCI(aq) \rightarrow H(aq) + CI(aq)$$

 An Arrhenius base is defined as any compound that dissociates in aqueous solution to form \_\_\_\_\_\_ ions.
 KOH (aq) → K\* (aq) + OH (aq)

NaOH (aq) 
$$\rightarrow$$
 Na (aq) + OH (aq)

Salts are compounds that dissociate in aqueous solution releasing neither
 H ions nor OH ions.

$$KCl(aq) \rightarrow K^*(aq) + Cl^*(aq)$$

Using the Arrhenius definition, classify the following examples as acids, bases, or salts:

HBr asset of KCI Service. beaute act d Mig(OHI)<sub>2</sub> HaPO<sub>4</sub> enert of HCIO act d HCL partie. Benzer KNIO<sub>2</sub> Al(OH)s HFO4 act of KC2H3O2 partit.

NaCl

bease

# **Note Taking Worksheet Acids Bases And Salts**

**JL Elias** 

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