

Chemical Formula

Talking about proportions, molecules and relative of atoms present in a compound.

Molecular $\text{C}_{10}\text{H}_{12}\text{O}_4$ (molecular)

Formula represents the number of atoms of each in a single molecule.

Empirical $\text{C}_5\text{H}_6\text{O}_2$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound.

Atomic Mass

The average mass of atoms present in a sample.

Mass

$\text{C}_{10}\text{H}_{12}\text{O}_4$ mass of one molecule

$\text{C}_5\text{H}_6\text{O}_2$ mass of one ion

Relative Mass

$\text{C}_{10}\text{H}_{12}\text{O}_4$ relative mass

$\text{C}_5\text{H}_6\text{O}_2$ relative mass

$\text{C}_{10}\text{H}_{12}\text{O}_4$

Note Taking Guide Physics 601

AW Rasmussen



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