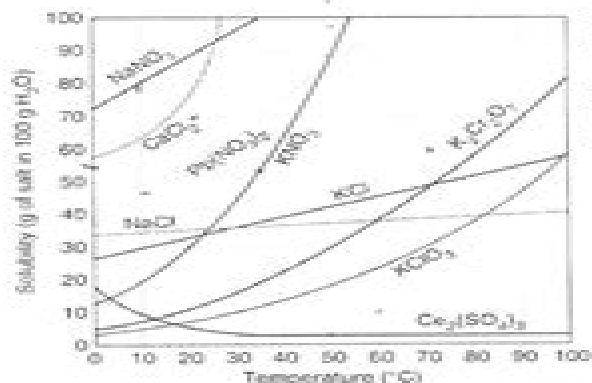


Worksheet: Solubility Graphs

Name

Key

Use the provided solubility graph to answer the following questions:



For questions 1 - 4 an amount of solute is given, and a temperature is stated. If all of the solute could be dissolved in 100 g of water at the given temperature, would the resulting solution be unsaturated, saturated, or supersaturated?

- 60 g KCl at 70 °C Super Saturated
- 10 g KClO₃ at 60 °C Unsaturated
- 80 g NaNO₃ at 10 °C Saturated
- 70 g CaCl₂ at 20 °C Unsaturated

For questions 5 - 8 a solute and temperature are given. Tell how many grams of each solute must be added to 100 g of water to form a saturated solution at the given temperature.

- Pb(NO₃)₂ at 10 °C 47g
- Ce₂(SO₄)₃ at 50 °C 3g
- NaCl at 20 °C 35g
- K₂Cr₂O₇ at 50 °C 30g

For questions 9 and 10 underline the solution that is more concentrated.

- At 10 °C: a saturated solution of KNO₃ or a saturated solution of CaCl₂
- At 50 °C: a saturated solution of KNO₃ or an unsaturated solution of NaNO₃ consisting of 90 g of the solute dissolved in 100 g of water.

For questions 11 - 12, show your work and circle your final answer.

- If 115 g KNO₃ are added to 100 g of water at 35 °C, how many grams do not dissolve? 55g = 60g will not dissolve.

- What mass of KCl would be needed to form a saturated solution if the KCl was dissolved in 200 g of water at 80 °C?

$$\frac{52g \text{ KCl}}{100g \text{ H}_2\text{O}} = \frac{104g \text{ KCl}}{200g \text{ H}_2\text{O}}$$

Same Ratio

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1. Understanding the eBook Progil Answers Solubility Chemistry
 - The Rise of Digital Reading Progil Answers Solubility Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Progil Answers Solubility Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Progil Answers Solubility Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Progil Answers Solubility Chemistry
 - Personalized Recommendations
 - Progil Answers Solubility Chemistry User Reviews and Ratings
 - Progil Answers Solubility Chemistry and Bestseller Lists
5. Accessing Progil Answers Solubility Chemistry Free and Paid eBooks
 - Progil Answers Solubility Chemistry Public Domain eBooks
 - Progil Answers Solubility Chemistry eBook Subscription Services
 - Progil Answers Solubility Chemistry Budget-Friendly Options

6. Navigating Progil Answers Solubility Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Progil Answers Solubility Chemistry Compatibility with Devices
 - Progil Answers Solubility Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Progil Answers Solubility Chemistry
 - Highlighting and Note-Taking Progil Answers Solubility Chemistry
 - Interactive Elements Progil Answers Solubility Chemistry
8. Staying Engaged with Progil Answers Solubility Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Progil Answers Solubility Chemistry
9. Balancing eBooks and Physical Books Progil Answers Solubility Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Progil Answers Solubility Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Progil Answers Solubility Chemistry
 - Setting Reading Goals Progil Answers Solubility Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Progil Answers Solubility Chemistry
 - Fact-Checking eBook Content of Progil Answers Solubility Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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