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 $\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{Pb} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2$
 $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$
 $\text{K}_2\text{SO}_4 + 6\text{KBr} \rightarrow \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2$
 $\text{CO}_2 + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5\text{O}_2$
 $3\text{CO}_2 + 4\text{H}_2\text{O} \rightarrow \text{C}_8\text{H}_{18} + 25\text{O}_2$
 $\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$

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 $\text{Na} + 3\text{PO}_4 + 3\text{KOH} + 3\text{NaOH} + 1\text{K} + 3\text{PO}_4 + 1\text{MgF}_2 + 1\text{Li} + 2\text{CO}_3 + 1\text{MgCO}_3 + 2\text{LiF} + 1\text{P}_4 + 3\text{O}_2 + 2\text{P}_2\text{O}_3 + 2\text{RbNO}_3 + 1\text{BeF}_2 + 1\text{Be} + \text{NO}_3 + 2\text{RbF}$

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 $\text{NaCl} + \text{F}_2 \rightarrow 2\text{NaF} + \text{Cl}_2$
 $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
 $\text{Pb} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{PbCl}_2$
 $2\text{Al} + 3\text{Br}_2 \rightarrow 2\text{AlBr}_3$
 $\text{K}_2\text{SO}_4 + 6\text{KBr} \rightarrow \text{Al}_2\text{SO}_4 + 3\text{CH}_4 + 2\text{O}_2$
 $\text{CO}_2 + 2\text{H}_2\text{O} \rightarrow \text{C}_3\text{H}_8 + 5\text{O}_2$
 $3\text{CO}_2 + 4\text{H}_2\text{O} \rightarrow \text{C}_8\text{H}_{18} + 25\text{O}_2$

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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $\text{HNO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{O} + \text{O}_2 + 3\text{H}_2\text{O}$
 $2\text{H}_2\text{O} + \text{O}_2 \rightarrow 4\text{H}_2\text{O}$
 $\text{RbH}_2\text{O} + \text{RbOH} + \text{H}_2\text{O} \rightarrow 5\text{Cl}_2 + 2\text{KBr} + \text{KCl} + \text{Br}_2 + 6\text{N}_2 + 2\text{H}_2 + \text{NH}_3 + 7\text{N}_2\text{F}_6 + 3\text{Ca} + \text{Nd} + \text{CaF}_2 + 8\text{CH}_4 + \text{H}_2\text{O} + \text{CO}_2 + 9\text{C}_2\text{H}_5\text{OH}$

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web balancing chemical equations ks3 gcse answers chemistrytutor me 2018 page 3 of 3 $38 \text{Fe} + 2 \text{O}_3 \rightarrow 3 \text{Co} + 2 \text{Fe}_3\text{O}_4$ $39 \text{H}_2 + 3 \text{PO}_4 \rightarrow 3 \text{BaO} + \text{Ba}_3\text{PO}_4$ $4 \text{H}_2 + 2 \text{O} \rightarrow 4 \text{H}_2\text{O}$ $40 \text{Mg} + \text{NH}_4\text{PO}_4 \rightarrow 4 \text{Mg}_3\text{P}_2 + 7 \text{N}_2 + 3 \text{H}_2\text{O}$ $41 \text{Al} + 3 \text{H}_2\text{SO}_4 \rightarrow 2 \text{Al}_2\text{O}_3 + 3 \text{H}_2 + 4 \text{H}_2\text{O}$ $10 \text{HCl} + \text{POCl}_3 \rightarrow 3 \text{H}_3\text{PO}_4 + 3 \text{H}_2\text{O} + 2 \text{H}_2\text{SO}_4 + 4 \text{HNO}_3$

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