

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

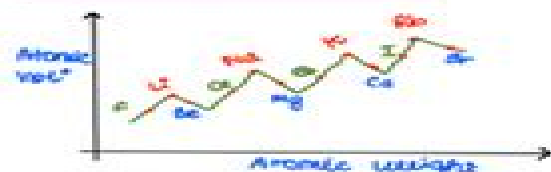
- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- based on Moseley law: $\sqrt{f} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e⁻ configuration.

Transition elements; they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G₁ and G₂ = ns¹⁻²

p-Block = G₁₃ - G₁₈ = ns² np¹⁻⁶

d-Block = G₃ - G₁₀ = (n-1)d¹⁻¹⁰ ns²

f-Block = lanthanoid = 2 = 58 - 71

G₃ and P₆

actinoids = 2 = 90 - 103

G₃ and P₆

Diagonal Relation;

Li → Be → B → similar
Na → Mg → Al → change ratio size

Nomenclature;

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	100 unnilpentium
2 → bi	7 → sept	
3 → tri	8 → oct	
4 → quad	9 → non	

EFFECTIVE NUCLEAR CHARGE

$Z_{\text{eff}} = Z - \sigma$

Trends - Period → $Z_{\text{eff}} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{\text{eff}} \uparrow \downarrow$

Shielding → s > p > d > f

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff}

IONISATION ENERGY

- Energy required to remove an electron from an isolated gaseous atom.



IE = IE₁ < IE₂ < IE₃ ...

Trends - $Z_{\text{eff}} \propto \text{IE}$

Period - $Z_{\text{eff}} \uparrow = \text{IE} \uparrow$

Group - $Z_{\text{eff}} \downarrow = \text{IE} \downarrow$

G₁₃ = G > Al > Ga > In

G₁₄ = C > Si > Ge > Pb > Sn

Period 2:-

Li < Be < B < C < N < O < F < Ne

ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covallent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends; Atomic Radius ∝ 1/Z_{eff}
Period → Z_{eff} ↑ = Atomic Radius ↓
Group → Z_{eff} ↓ = Atomic Radius ↑

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

- Energy required or released when an e⁻ is added to outermost orbit of an isolated gaseous atom. It is known as Electron gain enthalpy.



Note - In majority cases Δ_{eg}H₂ is -ve (exothermic)

Exception = N } Δ_{eg}H₂ = +ve
N } due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends; Δ_{eg}H ∝ Z_{eff}

Period - Z_{eff} ↑ → Δ_{eg}H ↑

Group - Z_{eff} ↓ → Δ_{eg}H ↓

G₁₃ = S > Se > Te > Po > O
→ due to small size Reason!
→ e⁻ repulsion.

Periodic Classification 3rd Chapter 11th Class

SJ Ball



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