

Problem Solving: Obtaining Empirical Formula from Experimental Data

PROCEDURE FOR

Obtaining an Empirical Formula from Experimental Data

1. Write down (or calculate) in grams the masses of each element present in a sample of the compound. If you are given mass percent composition, assume a 100-g sample and calculate the masses of each element from the given percentages.

2. Convert each of the masses in Step 1 to moles by using the appropriate molar mass for each element as a conversion factor.

3. Write down a preliminary formula for the compound using the number of moles of each element from Step 2 as subscripts.

4. Divide all the subscripts in the formula by the smallest subscript.

5. If the subscripts are not whole numbers, multiply all the subscripts by a small whole number (see table) to get whole-number subscripts.

Fractional Subscript	Multiply by This
0.50	2
0.25	4
0.33	3
0.40	5
0.20	2
0.60	5
0.75	4
0.80	5

EXAMPLE 5.15

Obtaining an Empirical Formula from Experimental Data

A laboratory analysis of aspirin determined the following mass percent composition:
C 60.00%
H 4.48%
O 35.52%
Find the empirical formula.

GIVEN: In a 100-g sample, 60.00 g C, 4.48 g H, 35.52 g O.

FOUN: empirical formula

$$60.00 \text{ g C} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 5.00 \text{ mol C}$$

$$4.48 \text{ g H} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 4.44 \text{ mol H}$$

$$35.52 \text{ g O} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 2.22 \text{ mol O}$$



The correct empirical formula is $\text{C}_9\text{H}_8\text{O}_4$.

FOR PRACTICE 5.15

Aspirin has the following mass percent composition:
C 75.00%, H 5.80%, O 19.20%
What is the empirical formula of aspirin?

Problem Solving Empirical Formula Answer Sheet

RJ Shavelson



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The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. *The Practice of Chemistry* First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools, and integrated media makes *The Practice of Chemistry* the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink. http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB

Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977

An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life. It provides a major overview of topics typically found in first year chemistry courses in the USA. The textbook is written in a conversational question-based format with a well-defined problem-solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box. Numerous examples are presented in every chapter to aid students and provide helpful self-learning tools. The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry.

Chemistry, Problem-Solving Worktext James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo.

Fireworks One of the ingredients is phosphorus, a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms. The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light.

Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball-shaped molecules often referred to as buckyballs. Diamond is not the most stable form of carbon and is created from other forms of carbon at high temperatures and pressures deep within the earth.

Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen, a reaction between molecules of C_2H_2 and O_2 to give CO_2 and H_2O .

Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane, each composed of three carbon and eight hydrogen atoms.

Stormy weather The evaporation of water serves to store energy provided by the sun. Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet.

Fire Engineering, 1927

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Chemistry Greg Curran, 2011 Covers all the topics in a typical one year high school chemistry curriculum

NEET 5000+ Chapter-wise SURESHOT Graded Problems in Physics, Chemistry & Biology 2nd Edition Disha Experts, 2019-11-14

Foundations of Chemistry Philippa B. Cranwell, Elizabeth M. Page, 2021-08-02 FOUNDATIONS OF CHEMISTRY A foundation level guide to chemistry for physical life sciences and engineering students Foundations of Chemistry An Introductory Course for Science Students fills a gap in the literature to provide a basic chemistry text aimed at physical sciences life sciences and engineering students The authors noted experts on the topic offer concise explanations of chemistry theory and the principles that are typically reviewed in most one year foundation chemistry courses and first year degree level chemistry courses for non chemists The authors also include illustrative examples and information on the most recent applications in the field Foundations of Chemistry is an important text that outlines the basic principles in each area of chemistry physical inorganic and organic building on prior knowledge to quickly expand and develop a student's knowledge and understanding Key features include Worked examples showcase core concepts and practice questions Margin comments signpost students to knowledge covered elsewhere and are used to highlight key learning objectives Chapter summaries list the main concepts and learning points

General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing, 1999

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Fluid Power Circuits and Controls John S. Cundiff, Michael F. Kocher, 2019-12-05 Fluid Power Circuits and Controls Fundamentals and Applications Second Edition is designed for a first course in fluid power for undergraduate engineering students After an introduction to the design and function of components students apply what they ve learned and consider how the component operating characteristics interact with the rest of the circuit The Second Edition offers many new worked examples and additional exercises and problems in each chapter Half of these new problems involve the basic analysis of specific elements and the rest are design oriented emphasizing the analysis of system performance The envisioned course does not require a controls course as a prerequisite however it does lay a foundation for understanding the extraordinary productivity and accuracy that can be achieved when control engineers and fluid power engineers work as a team on a fluid power design problem A complete solutions manual is available for qualified adopting instructors

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