

Electron Configuration Practice - Homework - **KEY**

In the space below, write the expanded electron configurations of the following elements:

- | | | |
|----|-----------|--|
| 1) | sodium | $1s^2 2s^2 2p^6 3s^1$ |
| 2) | potassium | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$ |
| 3) | chlorine | $1s^2 2s^2 2p^6 3s^2 3p^5$ |
| 4) | bromine | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$ |
| 5) | oxygen | $1s^2 2s^2 2p^4$ |

In the space below, write the abbreviated electron configurations of the following elements:

- | | | |
|-----|-----------|-------------------------|
| 6) | manganese | $[\text{Ne}] 3s^2$ |
| 7) | silver | $[\text{Kr}] 5s^2 4d^9$ |
| 8) | nitrogen | $[\text{He}] 2s^2 2p^3$ |
| 9) | sulfur | $[\text{Ne}] 3s^2 3p^4$ |
| 10) | argon | $[\text{Ne}] 3s^2 3p^6$ |

In the space below, write the orbital notation of the following elements:

- | | | | | | | | | | | | | | | | | | | | |
|---------------|-------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|----------------------|------------|--|
| 11) manganese | 12 e ⁻ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | | | | | | | | | | | | |
| | | 1s | 2s | 2p | 2p | 2p | 3s | | | | | | | | | | | | |
| 12) gallium | 31 e ⁻ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | \uparrow | |
| | | 1s | 2s | 2p | 2p | 2p | 3s | 3p | 3p | 3p | 4s | 3d | 3d | 3d | 3d | 3d | 3d | 4p | |
| 13) nitrogen | 7 e ⁻ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | \uparrow | \uparrow | \uparrow | | | | | | | | | | | | | |
| | | 1s | 2s | 2p | 2p | 2p | | | | | | | | | | | | | |
| 14) sulfur | 16 e ⁻ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | \uparrow | \uparrow | | | | | | | | | |
| | | 1s | 2s | 2p | 2p | 2p | 3s | 3p | 3p | 3p | | | | | | | | | |
| 15) argon | 18 e ⁻ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | $\uparrow\downarrow$ | | | | | | | | | |
| | | 1s | 2s | 2p | 2p | 2p | 3s | 3p | 3p | 3p | | | | | | | | | |

Determine what elements are denoted by the following electron configurations:

- | | | |
|-----|---|-------------|
| 16) | $1s^2 2s^2 2p^6 3s^2 3p^4$ | sulfur |
| 17) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ | rubidium |
| 18) | $[\text{Kr}] 5s^2 4d^{10} 5p^3$ | antimony |
| 19) | $[\text{Xe}] 6s^2 4f^{14} 5d^6$ | osmium |
| 20) | $[\text{Rn}] 7s^2 5f^{11}$ | einsteinium |

Determine which of the following electron configurations are not valid:

- | | | |
|-----|--|--|
| 21) | $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ | not valid (take a look at "4d") |
| 22) | $1s^2 2s^2 2p^6 3s^3 3d^5$ | not valid (3p comes after 3s) |
| 23) | $[\text{Ra}] 7s^2 5f^6$ | not valid (radium isn't a noble gas) |
| 24) | $[\text{Kr}] 5s^2 4d^{10} 5p^5$ | valid |
| 25) | $[\text{Xe}]$ | not valid (an element can't be its own electron configuration) |

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RS Peters



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