# Science 30 UNIT D: Energy and the Environment

# Harvesting Chemical Energy and Hess's Law Assignment

Please use the "Heat of Formation" values found in the Science 30 data booklet. Remember that there are different values for water vapour  $[H_2O(g)]$  and liquid water  $[H_2O(l)]$ .

- The following balanced chemical equations are for based on the combustions of various fuels.
   Calculate the heats of combustion for each reaction
  - a. C<sub>2</sub>H<sub>5</sub>OH (I) + 3 O<sub>2</sub> (g) → 2 CO<sub>2</sub> (g) + 3 H<sub>2</sub>O (g)

b.  $2 C_4H_{10}(g) + 13 O_2(g) \rightarrow 8 CO_2(g) + 10 H_2O(g)$ 

c.  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$ 

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