

SALT ANALYSIS



SCHEME OF ANALYSIS

EXPERIMENT	OBSERVATION	INFERENCE
PRELIMINARY EXAMINATIONS :		
1) Noted the colour	Colourless	Absence of $\text{Cu}^{2+}, \text{Fe}^{2+}, \text{Fe}^{3+}, \text{Mn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$
	Blue	Pre. of Cu^{2+}
	Brown	Pre. of Fe^{3+}
	Green	Pre. of Ni^{2+}
	Pale pink	Pre. of Mn^{2+}
	Pink	Pre. of Co^{2+}
2) Noted the solubility a) In water b) In dil.HCl	a) Soluble and no ppt with Na_2CO_3 b) Insoluble	a) Pre. of NH_4^+ b) Pre. of Pb^{2+}
3) Flame Test : A little of the salt is made into paste with conc.HCl. A glass rod is dipped in it and introduced into the flame.	a) Blue surrounded by green b) Brick red c) Crimson red d) Bluish white e) Grassy green on prolonged heating	a) Pre. of Cu^{2+} b) Pre. of Ca^{2+} c) Pre. of Sr^{2+} d) Pre. of Pb^{2+} e) Pre. of Ba^{2+}
4) Ash Test : A little of the salt is heated with con. HNO_3 and cobalt nitrate solution. A filter paper is dipped in it and burnt.	a) Green ash b) Blue tinted ash c) Pale pink ash	a) Pre. of Zn^{2+} b) Pre. of Al^{3+} c) Pre. of Mg^{2+}
SYSTEMATIC ANALYSIS OF ANIONS :		
1) $\text{dil.H}_2\text{SO}_4$ test : To the salt add $\text{dil.H}_2\text{SO}_4$, warm if required.	a) Evolution of Colourless, odourless gas with brisk effervescence, turns limewater milky. b) Smell of vinegar	a) Pre. of carbonates b) Pre. of acetate

Salt Analysis Manual For P G Course

Judd E. Hollander



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