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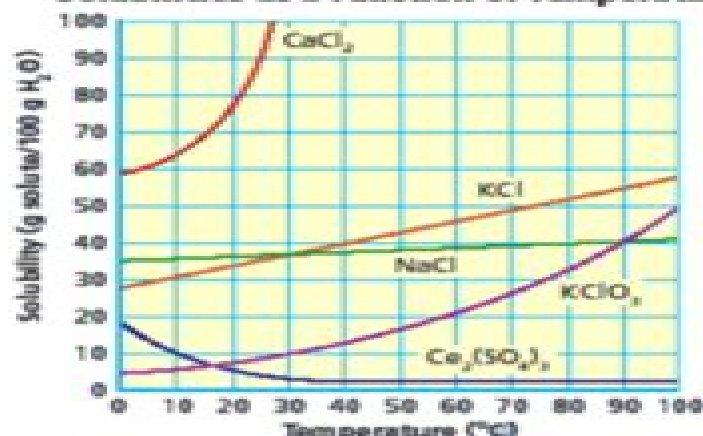
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Solubility Curve Worksheet 2



Directions: Use the graph below to answer the following questions.

Solubilities as a Function of Temperature



Types of Solutions

- Points that collectively make up the solubility curve (points **ON** the curve) represent _____ solutions.
- Points that are **BELOW** the curve represent _____ solutions.
- Points **ABOVE** the solubility curve represent _____ solutions, and the **difference** between the point above the curve and on the curve represents the amount of solute which will precipitate out.
- Say if the resulting solutions would be saturated, supersaturated or unsaturated.
 - 60 g of KCl at 70°C _____
 - 10 g of KClO₃ at 60°C _____
 - 70 g of CaCl₂ at 20°C _____

1. a. What is the solubility of calcium chloride (CaCl₂) at 5°C? _____
 b. What is the solubility of calcium chloride at 25°C? _____
 c. What is happening to the solubility and temperature for CaCl₂? Solubility is _____ as temperature is _____.
2. a. What is the solubility of cerium sulfate - Ce₂(SO₄)₃ at 10°C? _____
 b. What is the solubility of cerium sulfate at 50°C? _____
 c. What is happening to the solubility and temperature for Ce₂(SO₄)₃? Solubility is _____ as temperature is _____.
3. Which substance on the graph is **least** soluble at 10°C? _____
4. Which substance on the graph shows the **least** change in solubility from 0°C to 100°C? _____
5. Which solution is more concentrated:
 - a) At 10°C, a saturated solution of NaCl or a saturated solution of CaCl₂ (circle one)
 - b) At 55°C, a saturated solution of KClO₃ or a saturated solution of Ce₂(SO₄)₃ (circle one)
6. What is the mass of cerium sulfate that will dissolve in 50 g of water at 10 degrees Celsius?
7. What is the mass of potassium chloride that will dissolve in 250 g of water at 70 degrees Celsius?
8. At 90 degrees Celsius, 10 g of potassium chlorate is dissolved in 100 g of water. Is this solution saturated, unsaturated, or supersaturated? How do you know?
9. A saturated solution of potassium chlorate is dissolved in 100g of water. If the saturated solution is cooled from 90 degrees Celsius to 60 degrees Celsius, how many grams would crystallize out?

Practice Solubility Problems

S Ben Porath



Practice Solubility Problems:

This book delves into Practice Solubility Problems. Practice Solubility Problems is an essential topic that must be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Practice Solubility Problems, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Practice Solubility Problems
 - Chapter 2: Essential Elements of Practice Solubility Problems
 - Chapter 3: Practice Solubility Problems in Everyday Life
 - Chapter 4: Practice Solubility Problems in Specific Contexts
 - Chapter 5: Conclusion
2. In chapter 1, the author will provide an overview of Practice Solubility Problems. The first chapter will explore what Practice Solubility Problems is, why Practice Solubility Problems is vital, and how to effectively learn about Practice Solubility Problems.
3. In chapter 2, the author will delve into the foundational concepts of Practice Solubility Problems. This chapter will elucidate the essential principles that need to be understood to grasp Practice Solubility Problems in its entirety.
4. In chapter 3, the author will examine the practical applications of Practice Solubility Problems in daily life. The third chapter will showcase real-world examples of how Practice Solubility Problems can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Practice Solubility Problems in specific contexts. The fourth chapter will explore how Practice Solubility Problems is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Practice Solubility Problems. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Practice Solubility Problems.

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Table of Contents Practice Solubility Problems

1. Understanding the eBook Practice Solubility Problems
 - The Rise of Digital Reading Practice Solubility Problems
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Solubility Problems
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Solubility Problems
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Solubility Problems
 - Personalized Recommendations
 - Practice Solubility Problems User Reviews and Ratings
 - Practice Solubility Problems and Bestseller Lists
5. Accessing Practice Solubility Problems Free and Paid eBooks
 - Practice Solubility Problems Public Domain eBooks
 - Practice Solubility Problems eBook Subscription Services
 - Practice Solubility Problems Budget-Friendly Options
6. Navigating Practice Solubility Problems eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Solubility Problems Compatibility with Devices
 - Practice Solubility Problems Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Solubility Problems
 - Highlighting and Note-Taking Practice Solubility Problems
 - Interactive Elements Practice Solubility Problems
8. Staying Engaged with Practice Solubility Problems

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Solubility Problems
9. Balancing eBooks and Physical Books Practice Solubility Problems
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Solubility Problems
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Practice Solubility Problems
 - Setting Reading Goals Practice Solubility Problems
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Practice Solubility Problems
 - Fact-Checking eBook Content of Practice Solubility Problems
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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