

Periodic and Groups: Valence Electrons: Mastery

Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same column as alkali.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same column have similar properties.

Period	Group		Group					Group	
	1A	2A	3A	4A	5A	6A	7A	8A	18A
1	H 1H								He 2He
2	Li 3Li	Be 4Be							Ne 10Ne
3	Na 11Na	Mg 12Mg	Transition metals in between these rows						Ar 18Ar
4	K 19K	Ca 20Ca							Kr 36Kr

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Group 1A and Period 3?

What group and period is Chlorine in?

What element is in Group 2 and Period 18A? Chlorine is in Group 7A and Period 3A. The element is oxygen.

What element is in Group 2A and Period 2?

What element Magnesium?

Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that get involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Lithium 3



3 valence electrons

Hydrogen and Helium are both in Group 1A and Lithium is in Group 1A and Lithium has 3 valence electrons.

Group 1A, Lithium 3A and 5A have valence electrons. How many valence electrons will they have?

How about Lithium?

Group 18A: - These are noble gases that have a full shell of electrons. Elements are in Group 18 and they have 8 valence electrons. Helium has 2 valence electrons. "He 2 is, 2 electrons" Group 18A has a full shell of valence electrons normally. 18 valence electrons will have, gases or noble or inert or noble gases.

Periods And Groups Valence Electrons Masses Answers

Amar Nath Bhutani



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