

The Periodic Table of Elements

In 1869 Dimitri Mendeleev organized the elements in a table based on the properties of the elements.

Groups

↓ ↓
(Columns)

The elements in each group have the same number of valence electrons.

1 st Period → (1 shell) 	Group 1-2 (1-2 valence e ⁻)	1 H Hydrogen	Group 2-2 (2 valence electrons)	4 Be Beryllium	Group 3-2 (3 valence electrons)	5 B Boron	Group 4-2 (4 valence electrons)	6 C Carbon	Group 5-2 (5 valence electrons)	7 N Nitrogen	Group 6-2 (6 valence electrons)	8 O Oxygen	Group 7-2 (7 valence electrons)	9 F Fluorine	Group 8-2 (8 valence electrons)	2 He Helium	↳ Except Helium with 2 valence e ⁻
2 nd Period → (2 shells) 		3 Li Lithium														10 Ne Neon	
3 rd Period → (3 shells) 		11 Na Sodium		12 Mg Magnesium		13 Al Aluminum		14 Si Silicon		15 P Phosphorus		16 S Sulfur		17 Cl Chlorine		18 Ar Argon	
4 th Period → (4 shells) 		19 K Potassium		20 Ca Calcium		31 Ga Gallium		32 Ge Germanium		33 As Arsenic		34 Se Selenium		35 Br Bromine		36 Kr Krypton	

Periods →
(rows)

The elements in each period have the same number of shells.

Today, 114 elements are arranged by atomic number in the periodic table. The atomic number is based on the number of protons in an element.

Periodic Table Notes For Middle School

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LaVaque-Manty, Deborah Meizlish**

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