# **Balancing Chemical Reactions - Practice**

Balance the following reactions and identify what type of reactions they are.

Type of Reaction? Synthesis



Type of Reaction? Double Replacement

$$30 1 P_4 + 3 O_2 \rightarrow 2 P_2O_3$$

Type of Reaction? Synthesis

$$43$$
 2 RbNO<sub>3</sub> + 1 BeF<sub>2</sub>  $\rightarrow$  1 Be(NO<sub>3</sub>)<sub>2</sub> + 2 RbF

Type of Reaction? Double Replacement

s) 
$$2$$
 AgNO<sub>3</sub> +  $1$  Cu  $\rightarrow$   $1$  Cu(NO<sub>3</sub>)<sub>2</sub> +  $2$  Ag

Type of Reaction? Single Replacement

$$63 2 H_2O_2 \Rightarrow 2 H_2O + 1 O_2$$

Type of Reaction? Decomposition

7) 
$$2$$
 HCN +  $1$  Cu(CN)<sub>2</sub>

Type of Reaction? Double Replacement

$$m = 1$$
 GaF<sub>1</sub> + 3 Cs + 3 CsF + 1 Ga

Type of Reaction? Single Replacement

# **Practice With Identifying Chemical Reactions**

**Burdge** 

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