

Use the following information to answer questions 12-16.

trial	initial [A]	initial [B]	initial [C]	rate
#1	0.10	0.02	0.04	10 M/hr
#2	0.10	0.03	0.04	15 M/hr
#3	0.20	0.02	0.08	80 M/hr
#4	0.20	0.02	0.16	160 M/hr
#5	0.05	0.01	0.08	?

$$\Delta []^x = \Delta \text{Rate}$$

$$\left(\frac{1}{2}\right)^2 \left(\frac{1}{2}\right) \left(2\right) = \frac{1}{4} 10 \text{ M/hr}$$

12. What is the rate order for reactant C?

1st order

13. What is the rate order for reactant B?

1st order

14. What is the rate order for reactant A? (Hint: Use the information for the question 12 and trials #1 and #3)

2nd order

15. Calculate the rate constant and write out a rate law equation for the data given above.

$$\text{Rate} = k [A]^2 [B]^1 [C]^1$$

16. What would the rate be in trial #5?

2.5 M/hr Rate overall = 4

$$k = \frac{\text{Rate}}{[A]^2 [B]^1 [C]^1} \quad \frac{1}{\text{M}^3 \text{s}}$$

Practice Rate Law Problems Answers

Albert A Gayle



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