



NEW JERSEY CENTER FOR TEACHING & LEARNING

Aqueous Equilibria: Part 1 - Buffers & Titrations

PSI AP Chemistry

Name _____

Multiple Choice questions

Common-Ion Effect

1) The pH of a solution that contains 0.818 M acetic acid ($K_a = 1.77 \times 10^{-5}$) and 0.172 M sodium acetate is _____.

- A) 4.077 B) 5.434 C) 6.571 D) 8.370 E) 9.922

2) Calculate the pH of a solution prepared by dissolving 0.37 mol of formic acid (HCO_2H) and 0.23 mol of sodium formate (NaCO_2H) in water sufficient to yield 1.00 L of solution. The K_a of formic acid is 1.77×10^{-4} .

- A) 2.1 B) 10.46 C) 3.55 D) 2.31 E) 3.95

3) Calculate the pH of a solution prepared by dissolving 0.25 mol of benzoic acid ($\text{C}_6\text{H}_5\text{O}_2\text{H}$) and 0.15 of sodium benzoate ($\text{NaC}_6\text{H}_5\text{O}_2$) in water sufficient to yield 1.00 L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 4.41 B) 2.40 C) 3.97 D) 10.08 E) 4.19

4) Calculate the pH of a solution prepared by dissolving 0.15 mol of benzoic acid (HBz) and 0.30 mol of sodium benzoate in water sufficient to yield 1.00L of solution. The K_a of benzoic acid is 6.50×10^{-5} .

- A) 2.52 B) 3.89 C) 4.49 D) 10.16 E) 4.20

5) Calculate the pH of a solution prepared by dissolving 0.750 mol of NH_3 and 0.250 mol of NH_4Cl in water sufficient to yield 1.00 L of solution. The K_b of ammonia is 1.77×10^{-5} .

- A) 5.22 B) 4.27 C) 9.73 D) 8.78 E) 0.89

6) The pH of a solution prepared by dissolving 0.350 mol of solid methylamine hydrochloride ($\text{CH}_3\text{NH}_3\text{Cl}$) in 1.00L of 1.10M methylamine (CH_3NH_2) is _____. The K_b for methylamine is 4.4×10^{-4} . (Assume the final volume is 1.00 L.)

- A) 1.86 B) 2.66 C) 10.15 D) 11.14 E) 10.64

Buffers

7) What does a buffer do?

- a) Keeps the pH of a solution constant.

Practice Multiple Choice For Chem Aqueous Equilibria

Richard Bailey



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