

Mole Calculation Practice Worksheet

Answer the following questions:

- 1) How many moles are in 25 grams of water?

$$\frac{25 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.4 \text{ moles}$$

- 2) How many grams are in 4.5 moles of Li_2O ?

$$\frac{4.5 \text{ mol Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \times \frac{29.88 \text{ g Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} = 130 \text{ g}$$

- 3) How many molecules are in 23 moles of oxygen?

$$\frac{23 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol O}_2} = 1.4 \times 10^{25} \text{ molecules}$$

- 4) How many moles are in 3.4×10^{23} molecules of H_2SO_4 ?

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = 0.56 \text{ moles}$$

- 5) How many molecules are in 25 grams of NH_3 ?

$$\frac{25 \text{ g NH}_3}{17.03 \text{ g NH}_3} \times \frac{1 \text{ mol NH}_3}{1 \text{ mol NH}_3} \times \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1 \text{ mol NH}_3} = 8.8 \times 10^{23} \text{ molecules}$$

- 6) How many grams are in 8.2×10^{22} molecules of N_2I_4 ?

$$\frac{8.2 \times 10^{22} \text{ molecules N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \times \frac{1 \text{ mol N}_2\text{I}_4}{1 \text{ mol N}_2\text{I}_4} \times \frac{789.42 \text{ g N}_2\text{I}_4}{1 \text{ mol N}_2\text{I}_4} = 110 \text{ g}$$

Practice Equations In Moles

S Nieto

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Table of Contents Practice Equations In Moles

1. Understanding the eBook Practice Equations In Moles
 - The Rise of Digital Reading Practice Equations In Moles
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Equations In Moles
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Equations In Moles
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Equations In Moles
 - Personalized Recommendations
 - Practice Equations In Moles User Reviews and Ratings
 - Practice Equations In Moles and Bestseller Lists
5. Accessing Practice Equations In Moles Free and Paid eBooks
 - Practice Equations In Moles Public Domain eBooks
 - Practice Equations In Moles eBook Subscription Services
 - Practice Equations In Moles Budget-Friendly Options
6. Navigating Practice Equations In Moles eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Equations In Moles Compatibility with Devices
 - Practice Equations In Moles Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Equations In Moles
 - Highlighting and Note-Taking Practice Equations In Moles
 - Interactive Elements Practice Equations In Moles
8. Staying Engaged with Practice Equations In Moles

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Practice Equations In Moles
- 9. Balancing eBooks and Physical Books Practice Equations In Moles
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Equations In Moles
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Equations In Moles
 - Setting Reading Goals Practice Equations In Moles
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Equations In Moles
 - Fact-Checking eBook Content of Practice Equations In Moles
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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