

Annotated periodic table of the elements

1 H1.01Hydrogen												13 B10.81Boron						14 C12.01Carbon	15 N14.01Nitrogen	16 O16.00Oxygen	17 F18.99Fluorine	18 Ne20.18Neon
3 Li6.94Lithium		4 Be9.01Beryllium	Most common oxidation states: electronegativity (Pauli-Pauling)																			
11 Na22.99Sodium		12 Mg24.31Magnesium																				
19 K39.10Potassium		20 Ca40.08Calcium																				
37 Rb85.47Rubidium		38 Sr87.62Strontium																				
55 Cs132.91Cesium		56 Ba137.33Barium																				
87 Fr223Francium		88 Ra226Radium																				

CLASS COUNTERPARTS		
Metals		Nonmetals
Active	Pre-transition Alkali earth Alkaline	Cometoid
Transition		Intermediate
Post-transition		Metallous
Noble (Ru-Pt, Os-Au)		Noble gas

49 La 138.91 Lanthanum	50 Ce 140.12 Cerium	51 Pr 140.91 Praseodymium	52 Nd 144.24 Neodymium	53 Pm 144.91 Promethium	54 Sm 150.36 Samarium	55 Eu 151.96 Europium	56 Gd 157.25 Gadolinium	57 Tb 158.93 Terbium	58 Dy 162.50 Dysprosium	59 Ho 164.93 Holmium	60 Er 167.26 Erbium	61 Tm 168.93 Thulium	62 Yb 173.05 Ytterbium
63 Ac 227.07 Actinium	64 Th 232.04 Thorium	65 Pa 231.04 Protactinium	66 U 238.03 Uranium	67 Np 237.05 Neptunium	68 Pu 244.06 Plutonium	69 Am 243.06 Americium	70 Cm 247.07 Curium	71 Bk 247.07 Berkelium	72 Cf 251.08 Californium	73 Es 252.08 Einsteinium	74 Fm 257.10 Fermium	75 Md 258.10 Mendelevium	76 No 259 Nobelium

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Th onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metalloids to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

Group 1
This group is Sc-Y, and the alkaline earths are La-Be, and 4 to 10 (not 11-12). Trends going down Sc-Y and Be:

- La trend, resemble those of groups 1-2
- Lu trend, more resemble those of groups 4-10

Cometoid nonmetals
O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically cometoid nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Periodic Table Guide Answer Sheet

JS Bruner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the JS Bruner text.

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