

Boyle's Law

1. The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
2. Boyle's
3. Boyle's Law

Boyle's Theory and Boyle's Law

1. Boyle's of gases is caused by the elasticity of the gas molecules themselves that collide.
2. If the gas is compressible to small the volume is small which increases the pressure which is small.
3. Force exerts pressure on closed on the walls of the container.
4. Volume of the gas will decrease.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. If the gas is released, it maintains a volume of 900 ml. above the pressure inside of 800 Pa. What is the volume of the balloon at this altitude?

$$P_1 V_1 = P_2 V_2 \quad 800 \text{ Pa} \cdot 800 \text{ ml} = 800 \text{ Pa} \cdot V_2 \quad V_2 = 900 \text{ ml}$$

Ex. 2. A gas is pressure of 100 Pa and 100 ml. of gas has a volume of 100 ml. If the pressure is increased until the volume is 200 ml. what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad 100 \text{ Pa} \cdot 100 \text{ ml} = P_2 \cdot 200 \text{ ml} \quad P_2 = 50 \text{ Pa}$$

$$P_2 = 50 \text{ Pa}$$

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A Loxley



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