- What is the difference between electronegativity and ionization energy?
- 2. Why does fluorine have a higher ionization energy than iodine?
- 3. Why do elements in the same family generally have similar properties?
- 4. Which is the largest atom in Group 4a (14)?
- Which is the smallest atom in Group7a (17)?
- Which is the smallest atom in period 5?
- Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
- 8. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
- Arrange the following atoms in order of decreasing atomic radius. Na Al P Cl Mg.
- 10. For each of the following pairs, circle the element that is larger? e) Fe2" or Fe3"
  - a) N<sup>3</sup>- or F<sup>-1</sup> b) Mg2 or Ca2+
- 11. Circle the element in each pair has the larger radius?
  - b) O or O2 c) K\* or Cl d) P3 or S2 a) Mg or Mg2"
- 12. In each of the following pairs, circle the species with the higher first ionization energy: (e) B or Be
  - a) Li or Cs (b) Cl or Ar (c) Ca or Br (d) Na or Ne
- 13. In each of the following pairs, circle the species with the larger atomic radius:
  - (b) S or S2- (c) Cu+2 or Cu (d) He or H a) Mg or Ba (e) Na or Cl
- 14. Circle the best choice in each list:
  - a) highest first ionization energy: C, N, Si
  - b) highest electronegativity: As, Sn, S
  - c) largest radius: S2-, C1-, C1
- Order the following groups from largest to smallest radii.
  - a) Ar. Cl. K. S<sup>2</sup>
  - b) C, Al, F, Si
  - c) Na. Mg. Ar. P.
  - d) F<sup>1</sup>, Ba<sup>2+</sup>, Cs<sup>+1</sup>, F<sup>-1</sup>

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#### **Table of Contents Periodic Trends Extensoin Questions Worksheet**

- 1. Understanding the eBook Periodic Trends Extensoin Questions Worksheet
  - The Rise of Digital Reading Periodic Trends Extensoin Questions Worksheet
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Extensoin Questions Worksheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Extensoin Questions Worksheet
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Extensoin Questions Worksheet
  - Personalized Recommendations
  - Periodic Trends Extensoin Questions Worksheet User Reviews and Ratings
  - Periodic Trends Extensoin Questions Worksheet and Bestseller Lists
- 5. Accessing Periodic Trends Extensoin Questions Worksheet Free and Paid eBooks

- Periodic Trends Extensoin Ouestions Worksheet Public Domain eBooks
- Periodic Trends Extensoin Questions Worksheet eBook Subscription Services
- Periodic Trends Extensoin Questions Worksheet Budget-Friendly Options
- 6. Navigating Periodic Trends Extensoin Questions Worksheet eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodic Trends Extensoin Questions Worksheet Compatibility with Devices
  - Periodic Trends Extensoin Questions Worksheet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Extensoin Questions Worksheet
  - Highlighting and Note-Taking Periodic Trends Extensoin Questions Worksheet
  - Interactive Elements Periodic Trends Extensoin Questions Worksheet
- 8. Staying Engaged with Periodic Trends Extensoin Questions Worksheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Extensoin Questions Worksheet
- 9. Balancing eBooks and Physical Books Periodic Trends Extensoin Questions Worksheet
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Periodic Trends Extensoin Questions Worksheet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Extensoin Questions Worksheet
  - Setting Reading Goals Periodic Trends Extensoin Questions Worksheet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Extensoin Questions Worksheet
  - Fact-Checking eBook Content of Periodic Trends Extensoin Questions Worksheet
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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