THE NATURE OF GASES

Section Review

Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- kimetic energy
- kimetic theory
- gas pressure

- Wasculline.
- atmospheric pressure
- Inarproperter

- pascal (Pa)
- standard atmosphere (atm):

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number,

The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly ______, that the particles of gas are relatively 3 move 4 of each other. and are in constant ___5__ motion. The ___6__ between particles are perfectly elastic so that the total ____7___ remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure _____8 ___ pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- 11. Atmospheric pressure is 760 mm Hg.
- 12. The SI unit of pressure is the pascal.

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web smart recovery a non profit corporation was originally named the rational recovery self help network and was affiliated with rational recovery systems a for profit corporation owned by jack trimpey in 1994 the non profit changed its name to smart recovery and ended all affiliation with trimpey

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