



Chemical Changes

Read the passage below and answer the following questions:



1. What are the products of this chemical reaction?

Mercury and oxygen

2. What is the reactant? Mercury oxide

3. What is the physical state of the reactant and the products?

The reactant is a solid, the products are a liquid and a gas.

4. According to the law of conservation of mass, if 100g of mercury oxide is heated, how much mercury and oxygen will be produced?

100g



5. What products will be produced in this reaction? Mercury and oxygen

6. What mass of reactants was used? 100g of mercury oxide

7. What are the physical states of the reactant and the products?

Reactant is solid.

8. What is the physical state of the product? Mercury is a liquid and oxygen is a gas.

9. What is the mass of the products? 100g of mercury and oxygen.

Mass is conserved.

10. How many grams of product are produced? 100g

Read the passage below and answer the following questions:

11. How many molecules of H_2O will be produced if 100g of hydrogen gas reacts with 100g of oxygen gas? 100g of H_2O will be produced.

Mass is conserved.

Section One Reinforcement Chemical Changes

Worksheet Answers

**Holt Rinehart & Winston, Holt,
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