

Questions 13–15 refer to the information below.

A 0.50 g sample of $\text{Mg}(s)$ was placed in a solution of $\text{HCl}(aq)$, where it reacted completely.

13. Which of the following equations best represents the reacting species in the reaction described above?
- (A) $\text{Mg}(s) + 2 \text{H}^+(aq) \rightarrow \text{Mg}^{2+}(aq) + \text{H}_2(g)$
 - (B) $\text{Mg}(s) + \text{HCl}(aq) \rightarrow \text{MgCl}_2(aq) + \text{H}_2(g)$
 - (C) $\text{Mg}(s) + 2 \text{HCl}(aq) \rightarrow \text{MgCl}_2(s) + 2 \text{H}^+(aq)$
 - (D) $\text{Mg}(s) + 2 \text{HCl}(aq) \rightarrow \text{Mg}^{2+}(aq) + \text{Cl}_2(g) + \text{H}_2(g)$
14. In another experiment, a small piece of $\text{Mg}(s)$ is weighed, then placed in a flask containing excess 1 *M* $\text{HCl}(aq)$. The student wants to determine number of moles of gas produced. Which of the following is the best way to conduct the experiment for accurate data collection?
- (A) Conducting the experiment at different temperatures to see which generates the most gas
 - (B) Completing the entire reaction in a large Erlenmeyer flask of known volume to measure the volume of the gas collected
 - (C) Collecting the gas in a eudiometer tube and measuring the volume of the gas collected
 - (D) Conducting the reaction in a graduated cylinder and measuring the volume of the gas collected
15. In a third experiment, 0.10 g samples of $\text{Mg}(s)$ are placed in excess $\text{HCl}(aq)$ of various concentrations: 0.050 *M*, 0.10 *M*, 0.25 *M*, and 0.50 *M*. The reactions are run in successive order from 0.050 *M* to 0.50 *M*, and the time required for each reaction to go to completion is recorded. As the concentration of $\text{HCl}(aq)$ increases from 0.050 *M* to 0.50 *M*, which of the following is the expected result?
- (A) The reaction time increases, and the rate of the reaction decreases.
 - (B) The reaction time decreases, and the rate of the reaction increases.
 - (C) Both the reaction time and the rate of the reaction increase.
 - (D) Both the reaction time and the rate of the reaction decrease.

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TD Snyder



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