

8.2

THE NATURE OF COVALENT BONDING

KEY

Section Review

Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

- stable
- covalent
- shared
- single
- unshared pairs
- double/triple
- coordinate covalent bond
- energy
- bond dissociation energy
- resonance structure

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KJ Lindholm-Leary



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