

Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



| Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-------|-------------------|-------------------------------------|----------------------------|
| Li    | 1 ve              | transfers 1ve <sup>-</sup> to F     | $\text{Li}^{+}$            |
| F     | 7 ve              | accepts 1ve <sup>-</sup> from Li    | $\text{F}^{-}$             |

| Reactions                                                             | Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-----------------------------------------------------------------------|-------|-------------------|-------------------------------------|----------------------------|
| 1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$                    | Li    | 1 ve              | donates 1e <sup>-</sup>             | $\text{Li}^{+1}$           |
| $\text{Li} + \text{Cl} \rightarrow [\text{Li}]^{+1} [\text{Cl}]^{-1}$ | Cl    | 1 ve              | accepts 1e <sup>-</sup>             | $\text{Cl}^{-1}$           |
| 2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$                      | Ca    | 2 ve              | donates 2e <sup>-</sup>             | $\text{Ca}^{+2}$           |
| $\text{Ca} + \text{O} \rightarrow [\text{Ca}]^{+2} [\text{O}]^{2-}$   | O     | 6 ve              | accepts 2e <sup>-</sup>             | $\text{O}^{2-}$            |
| 3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$                    | Be    | 2 ve              | donates 2e <sup>-</sup> to each F   | $\text{Be}^{+2}$           |
| $\text{Be} + \text{F} \rightarrow [\text{Be}]^{+2} [\text{F}]^{-1}$   | F     | 7 ve              | each F accepts 1ve each             | $\text{F}^{-1}$            |
| 4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$                      | Mg    | 2 ve              | donates 2e <sup>-</sup> to S        | $\text{Mg}^{+2}$           |
| $\text{Mg} + \text{S} \rightarrow [\text{Mg}]^{+2} [\text{S}]^{2-}$   | S     | 6 ve              | accepts 2e <sup>-</sup>             | $\text{S}^{2-}$            |
| 5) $\text{K} + \text{F} \Rightarrow \text{KF}$                        | K     | 1 ve              | donates 1ve <sup>-</sup>            | $\text{K}^{+1}$            |
| $\text{K} + \text{F} \rightarrow [\text{K}]^{+1} [\text{F}]^{-1}$     | F     | 7 ve              | accepts 1ve <sup>-</sup>            | $\text{F}^{-1}$            |

## Section 6 2 Ionic Bonding Worksheet Answer Key

**C Cleary**



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