

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Section 5 3 Electron Configuration Answers

Sebastian Brünink



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Table of Contents Section 5 3 Electron Configuration Answers

1. Understanding the eBook Section 5 3 Electron Configuration Answers
 - The Rise of Digital Reading Section 5 3 Electron Configuration Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 5 3 Electron Configuration Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 5 3 Electron Configuration Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 5 3 Electron Configuration Answers

- Personalized Recommendations
- Section 5 3 Electron Configuration Answers User Reviews and Ratings
- Section 5 3 Electron Configuration Answers and Bestseller Lists
- 5. Accessing Section 5 3 Electron Configuration Answers Free and Paid eBooks
 - Section 5 3 Electron Configuration Answers Public Domain eBooks
 - Section 5 3 Electron Configuration Answers eBook Subscription Services
 - Section 5 3 Electron Configuration Answers Budget-Friendly Options
- 6. Navigating Section 5 3 Electron Configuration Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 5 3 Electron Configuration Answers Compatibility with Devices
 - Section 5 3 Electron Configuration Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 5 3 Electron Configuration Answers
 - Highlighting and Note-Taking Section 5 3 Electron Configuration Answers
 - Interactive Elements Section 5 3 Electron Configuration Answers
- 8. Staying Engaged with Section 5 3 Electron Configuration Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 5 3 Electron Configuration Answers
- 9. Balancing eBooks and Physical Books Section 5 3 Electron Configuration Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 5 3 Electron Configuration Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 5 3 Electron Configuration Answers
 - Setting Reading Goals Section 5 3 Electron Configuration Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 5 3 Electron Configuration Answers

- Fact-Checking eBook Content of Section 5 3 Electron Configuration Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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