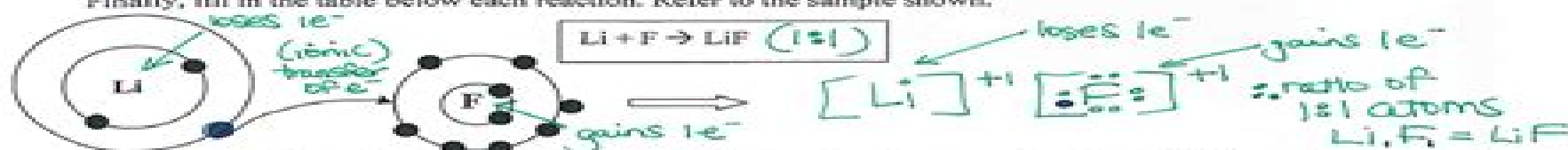


Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
Li	1 ve	transfers $1e^-$ to F	Li^+
F	7 ve	accepts $1e^-$ from Li	F^-

Reactions	Atoms	Valence electrons	Electron transfer from/to each atom	Ions formed in the product
1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$ $\text{Li} + \cdot\ddot{\text{Cl}}\cdot \rightarrow [\text{Li}]^+ [\ddot{\text{Cl}}:]^-$	Li	1 ve	donates $1e^-$	Li^+
	Cl	1 ve	accepts $1e^-$	Cl^-
2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$ $\text{Ca} + \cdot\ddot{\text{O}}\cdot \rightarrow [\text{Ca}]^{+2} [\ddot{\text{O}}:]^{2-}$	Ca	2 ve	donates $2e^-$	Ca^{+2}
	O	6 ve	accepts $2e^-$	O^{2-}
3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$ $\text{Be} + \cdot\ddot{\text{F}}\cdot \rightarrow [\text{Be}]^{+2} [\ddot{\text{F}}:]^{-1}$	Be	2 ve	donates $2e^-$ to each F	Be^{+2}
	F	7 ve	each F accepts $1e^-$ each	F^-
4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$ $\text{Mg} + \cdot\ddot{\text{S}}\cdot \rightarrow [\text{Mg}]^{+2} [\ddot{\text{S}}:]^{2-}$	Mg	2 ve	donates $2e^-$ to S	Mg^{+2}
	S	6 ve	accepts $2e^-$	S^{2-}
5) $\text{K} + \text{F} \Rightarrow \text{KF}$ $\text{K} + \cdot\ddot{\text{F}}\cdot \rightarrow [\text{K}]^+ [\ddot{\text{F}}:]^{-1}$	K	1 ve	donates $1e^-$	K^+
	F	7 ve	accepts $1e^-$	F^-

Section 6 Ionic Bonding Worksheet Answer Key

Keith Taber



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