

Electron Configurations - KEY

Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).

- 1) sodium $1s^2 2s^2 2p^6 3s^1$
- 2) iron $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$
also might be written (due to the close energy of d and f and similarities of elements versus predictable)
memorized config: $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt $[Ar] 4s^2 3d^7$
- 7) silver $[Kr] 5s^2 4d^9$
- 8) tellurium $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium $[Rn] 7s^2$
- 10) lawrencium $[Rn] 7s^2 5f^{14} 6d^1$
also might be written (due to the close energy of d and f and similarities of elements versus predictable)
memorized config: $[Rn] 7s^2 6d^1 5f^{14}$
- 11) $1s^2 2s^2 2p^6 3s^2 3p^4$ sulfur
- 12) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$ rubidium
- 13) $[Kr] 5s^2 4d^{10} 5p^3$ antimony
- 14) $[Xe] 6s^2 4f^{14} 5d^6$ osmium
- 15) $[Rn] 7s^2 5f^{11}$ einsteinium note: no d's, so the p. table referenced has Ac as 5f
- 16) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$ not valid (take a look at "4d"; must be 3 d not 4!)
- 17) $1s^2 2s^2 2p^6 3s^3 3d^5$ not valid (3p comes after 3s); also no 3 electrons in s
- 18) $[Ra] 7s^2 5f^8$ not valid (radium isn't a noble gas)
- 19) $[Kr] 5s^2 4d^{10} 5p^3$ valid
- 20) $[Xe]$ not valid (an element can't be its own electron configuration)

Section 5 3 Electron Configurations Answers

Marco Cascella



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Table of Contents Section 5 3 Electron Configurations Answers

1. Understanding the eBook Section 5 3 Electron Configurations Answers
 - The Rise of Digital Reading Section 5 3 Electron Configurations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 5 3 Electron Configurations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 5 3 Electron Configurations Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 5 3 Electron Configurations Answers
 - Personalized Recommendations
 - Section 5 3 Electron Configurations Answers User Reviews and Ratings
 - Section 5 3 Electron Configurations Answers and Bestseller Lists
5. Accessing Section 5 3 Electron Configurations Answers Free and Paid eBooks
 - Section 5 3 Electron Configurations Answers Public Domain eBooks
 - Section 5 3 Electron Configurations Answers eBook Subscription Services
 - Section 5 3 Electron Configurations Answers Budget-Friendly Options

6. Navigating Section 5 3 Electron Configurations Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 5 3 Electron Configurations Answers Compatibility with Devices
 - Section 5 3 Electron Configurations Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 5 3 Electron Configurations Answers
 - Highlighting and Note-Taking Section 5 3 Electron Configurations Answers
 - Interactive Elements Section 5 3 Electron Configurations Answers
8. Staying Engaged with Section 5 3 Electron Configurations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 5 3 Electron Configurations Answers
9. Balancing eBooks and Physical Books Section 5 3 Electron Configurations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 5 3 Electron Configurations Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 5 3 Electron Configurations Answers
 - Setting Reading Goals Section 5 3 Electron Configurations Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 5 3 Electron Configurations Answers
 - Fact-Checking eBook Content of Section 5 3 Electron Configurations Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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