6.00 It is often useful to write the equations for the reactions described:

 CH√(g) + 2Oy(g) → COy(g) + 2H√O(§) 	AH" = -890, NJ
2. Ctro + Ovigs -+ COvigs	JUNE 2014 NJ
3. Hules + 1000-last -+ Hu003	AA4" + -286 NJ

The reaction for which the heat of reaction is required is the formation of methane from its constituted elements.

The AV for this reaction can be calculated by summing the AV's for a series of reactions which together add up to the same ownall reaction (Hessi's Law of Heat Summation).

These equations add up to the desired equation, so the AH values are added:

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactions in the balanced equation:

There are two motes of benzene in the equation, so the heat released for one mote is 1/2 this

ampunt, 3139 kulmet.

6.130 a. Heat lost = 30, o x 0.126 J/s K x 66°C = 260 J

q(se melting) + q(water warming) + -q(water pooling)
 6.0 g x 334 pig + 5.0 g x 4.184 pig*C (T₁ - 0°C) + -100 g x 4.184 pig*C (T₁ - 70°C)
 1670 p + 20.92 T₁ + -418.4 T₁ + 29.286 p
 439.32 T₁ + 27.618
 T₁ + 62.8 + 63°C

6.14) a. q (water) - -q (coins)

ex ma AT = -ex ma AT 4.184 Jig*C = 200, g x 5.9°C = -ex 113.4 g = (-74.1°C). Solving for a gives: 0.587 Jig*C

No. the specific heat capacity of pure gold is 0.129 Jig*C.

8.15) J.H = 4.16 J/g K + 22.4 K + 70, g = 6550 J

For one mole of ammornium nitrate, 5W = 6550 x 60.05/20 = 26 kultimox.

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