

11-1 Practice Problems

1. Lead will react with Hydrochloric acid to produce Lead (II) Chloride and Hydrogen gas. How many moles of Hydrochloric acid are needed to completely react with 0.36-mol of lead?

2. How many moles of HNO_3 will be produced when 0.51-mol of N_2O_5 reacts according to the following equation?
 $\text{N}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$

3. Iron will react with Oxygen Gas to produce Fe_2O_3 . How many moles of Fe_2O_3 will be produced if 0.18-mol of Fe reacts?

4. How many moles of NaBr will be produced when 0.69-mol of Bromine reacts according to the following equation?
 $\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$

5. Phosphorus will react with Bromine to produce Phosphorus Tribromide. How many moles of Phosphorus Tribromide will be produced if 0.78-mol of Bromine is reacted?

6. Carbon will react with Zinc Oxide to produce Zinc and Carbon Dioxide. How many moles of carbon dioxide will be produced if 0.38-mol of ZnO is completely reacted?

7. How many moles of oxygen will be needed to react with 0.38-mol of C_3H_8 according to the following equation?
 $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

8. Nitrogen gas can react with Hydrogen gas to produce ammonia. How many moles of Nitrogen will be needed to produce 0.48-mol of NH_3 ?

9. How many moles of hydrogen will be produced if 0.44-mol of CaH_2 reacts according to the following equation?
 $\text{CaH}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$

10. How many moles of water will be produced if 2.35-mol of oxygen reacts according to the following equation?
 $2\text{C}_6\text{H}_6 + 15\text{O}_2 \rightarrow 12\text{CO}_2 + 6\text{H}_2\text{O}$

Prentice Hall Chemistry Solutions Practice Problem Answer

James F. Shackelford



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simplest mathematics along with numerous numerical examples it accurately and rigorously covers physical and chemical equilibrium in depth and detail It continues to cover the topics found in the first edition however numerous updates have been made including Changes in naming and notation the first edition used the traditional names for the Gibbs Free Energy and for Partial Molal Properties this edition uses the more popular Gibbs Energy and Partial Molar Properties changes in symbols the first edition used the Lewis Randal fugacity rule and the popular symbol for the same quantity this edition only uses the popular notation and new problems have been added to the text Finally the second edition includes an appendix about the Bridgman table and its use

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