

Chemistry - Review Chapters 4-5: 6

Directions: Answer each of the following questions. You need not use complete sentences.

1. By what property did Mendeleev arrange the elements?

ATOMIC MASS

2. By what property did Moseley suggest that the periodic table be arranged?

ATOMIC NUMBER

3. What is a period? How many are there in the periodic table?

A HORIZONTAL ROW IN THE PERIODIC TABLE; 7

4. What is a group (also called a family)? How many are there in the periodic table?

A VERTICAL COLUMN IN THE PERIODIC TABLE; 18

5. State the number of valence electrons in an atom of:

a. sulfur	b. calcium	c. chlorine	d. arsenic
6	2	7	5

6. Give the names and chemical symbols for the elements that correspond to these atomic numbers:

a. 10	b. 18	c. 36	d. 90
Ne, NEON	Ar, ARGON	Kr, KRYPTON	Th, THORIUM

7. List, by number, both the period and group of each of these elements.

	<u>Symbol</u>	<u>Period</u>	<u>Group</u>
a. beryllium	Be	2	2
b. iron	Fe	4	8
c. lead	Pb	6	14

8. Which of the following pairs of elements belong to the same period?

a. Na and Cl	b. Na and Li	c. Na and Cu	d. Na and Ne
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9. Which of the following pairs of elements belong to the same group?

a. H and He	b. Li and Be	c. C and Pb	d. Ga and Ge
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10. How does an element's period number relate to the number of the energy level of its valence electrons?

PERIOD NUMBER = ENERGY LEVEL OF VALENCE ELECTRONS

11. What are the transition elements?

GROUPS 3-12

12. In what type of orbitals are the actinide and lanthanide electrons found?

f ORBITALS

Prentice Hall Chemistry Section 6 Answer Key

S Marginson



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