the commence of the second of the commence of	Name:
u must show all work f	for credit.
e Date:	(No late packets will be accepted)
I. Balance Equa	tions
Balance the equations	\$
<ol> <li>Mg + ¼ H<sub>2</sub>O →</li> </ol>	$Mg(OH)_2 + H_2$
2. Mg + 2NaBr →	MgBr <sub>2</sub> + ZNa
3. C <sub>3</sub> H <sub>8</sub> + 5 O <sub>2</sub> → 3	CO <sub>2</sub> + 4 H <sub>2</sub> O
4. 4 P + (O <sub>2</sub> → 2 P <sub>2</sub> ( 5. Fe <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> + 6 KC	O5 OH →3 K₂SO₄ + 2 Fe(OH)₃
6. Fe <sub>2</sub> O <sub>3</sub> + 3 H <sub>2</sub> →	Fe + T H <sub>2</sub> O
7. How many moles	of hydrogen gas react with 1 mole of Fe <sub>2</sub> O <sub>3</sub> in reaction #6?
rite the helenced equa	tion for problems # 8 and 9:
8. Hydrogen gas and	I nitrogen monoxide combine to form water and nitrogen gas
o. Ilyanogon gas an	osphate and calcium chloride are mixed they form calcium phosphate and
9. When sodium pho	sphate and calcium chloride are mixed they form calcium phosphate and
sodium chloride	2 N- 5704 13 Cacle - Cas(PO)2 + ENac
	g balanced equation, prove (or explain) the Law of Conservation of Mass.
$N_2 + 3H_2 \rightarrow 2NF$	
N2 + 3H2 -> 2N1	See no pl gets created or lost
	see no plante created at lost
The state of the s	
11. Using the balance	ed equation shown in problem #10, explain what the coefficients represent
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	ed equation shown in problem #10, explain what the coefficients represent
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12. Using the balance	ed equation shown in problem #10, explain what the coefficients represent ed equation shown in problem #10, list the products of the reaction.
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