

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the word.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at low pressures.
5. A constant of nature that has values of 8.314 J/mol·K or 0.08206 L·atm/mol·K.
6. The a parameter of gas constant equations for determining the gas constant from experimental data.
7. An experimental gas constant derived from the ideal gas law at low pressures.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. A constant of nature that has values of 8.314 J/mol·K or 0.08206 L·atm/mol·K.
10. Ratio of volume to mass.

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. universal gas constant, pressure
- h. volume
- i. compressibility
- j. compressibility coefficient

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. According to Graham's law, the rate of effusion of a gas is proportional to the square root of its molar mass. Which of the following gases effuses most rapidly?
a. the slowest of the four
b. the one having the least mass
c. the slowest of the four
d. the one having the most mass

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L Darling-Hammond



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