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## 17

## THERMOCHEMISTRY

**SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK (pages 505–510)**

*This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat.*

**► Energy Transformations (page 505)**

- What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions? thermochemistry
- Where the use of energy is concerned (in a scientific sense), when is work done?  
Work is done when a force is used to move an object.
- Circle the letter next to each sentence that is true about energy:
  - ☒ Energy is the capacity for doing work or supplying heat.
  - ☒ Energy is detected only because of its effects.
  - ☐ Heat is energy that transfers from one object to another because they are at the same temperature.
  - ☒ Gasoline contains a significant amount of chemical potential energy.
- Circle the letter next to each sentence that is true about heat.
  - ☒ One effect of adding heat to a substance is an increase in the temperature of that substance.
  - ☐ Heat always flows from a cooler object to a warmer object.
  - ☒ If two objects remain in contact, heat will flow from the warmer object to the cooler object until the temperature of both objects is the same.

**► Exothermic and Endothermic Processes (pages 506–507)**

- What can be considered the “system” and what are the “surroundings” when studying a mixture of chemicals undergoing a reaction? Write your answers where indicated below.

System: The mixture of chemicals itself is considered the system.

Surroundings: Everything but the mixture of chemicals is the surroundings, but practically speaking, the immediate vicinity of the system.

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## **Table of Contents Prentice Hall Chemistry Work Answers Chapter 1**

1. Understanding the eBook Prentice Hall Chemistry Work Answers Chapter 1
  - The Rise of Digital Reading Prentice Hall Chemistry Work Answers Chapter 1
  - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Work Answers Chapter 1
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Prentice Hall Chemistry Work Answers Chapter 1
  - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Work Answers Chapter 1
  - Personalized Recommendations
  - Prentice Hall Chemistry Work Answers Chapter 1 User Reviews and Ratings
  - Prentice Hall Chemistry Work Answers Chapter 1 and Bestseller Lists
5. Accessing Prentice Hall Chemistry Work Answers Chapter 1 Free and Paid eBooks
  - Prentice Hall Chemistry Work Answers Chapter 1 Public Domain eBooks
  - Prentice Hall Chemistry Work Answers Chapter 1 eBook Subscription Services
  - Prentice Hall Chemistry Work Answers Chapter 1 Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Work Answers Chapter 1 eBook Formats

- ePub, PDF, MOBI, and More
- Prentice Hall Chemistry Work Answers Chapter 1 Compatibility with Devices
- Prentice Hall Chemistry Work Answers Chapter 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Work Answers Chapter 1
  - Highlighting and Note-Taking Prentice Hall Chemistry Work Answers Chapter 1
  - Interactive Elements Prentice Hall Chemistry Work Answers Chapter 1
- 8. Staying Engaged with Prentice Hall Chemistry Work Answers Chapter 1
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Prentice Hall Chemistry Work Answers Chapter 1
- 9. Balancing eBooks and Physical Books Prentice Hall Chemistry Work Answers Chapter 1
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Prentice Hall Chemistry Work Answers Chapter 1
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Prentice Hall Chemistry Work Answers Chapter 1
  - Setting Reading Goals Prentice Hall Chemistry Work Answers Chapter 1
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Prentice Hall Chemistry Work Answers Chapter 1
  - Fact-Checking eBook Content of Prentice Hall Chemistry Work Answers Chapter 1
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
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