

## Study Guide

## Key Concepts

## 14.1 Properties of Gases

- Gases are easily compressed because of the space between particles in a gas.
- The amount of gas ( $n$ ), volume ( $V$ ), and temperature ( $T$ ) are factors that affect gas pressure ( $P$ ).

## 14.2 The Gas Laws

- As the pressure of a gas increases, the volume decreases, if the temperature is constant.
- As the temperature of an enclosed gas increases, the volume increases, if the pressure is constant.
- As the temperature of an enclosed gas increases, the pressure increases, if the volume is constant.

- The combined gas law allows you to do calculations for situations in which only the amount of gas is constant.

## 14.3 Ideal Gases

- To calculate the number of moles of a contained gas requires an expression that contains the variable  $n$ .
- Real gases differ most from an ideal gas at low temperatures and high pressures.

## 14.4 Gases: Mixtures and Movements

- In a mixture of gases, the total pressure is the sum of the partial pressures of the gases.
- Gases of lower molar mass diffuse and effuse faster than gases of higher molar mass.

## Vocabulary

- Boyle's law (p. 418)
- Charles's law (p. 420)
- combined gas law (p. 424)
- compressibility (p. 413)
- Dalton's law of partial pressures (p. 432)
- diffusion (p. 435)
- effusion (p. 435)
- Gay-Lussac's law (p. 422)
- Graham's law of effusion (p. 435)
- ideal gas constant (p. 426)
- ideal gas law (p. 426)
- partial pressure (p. 432)

## Key Equations

- Boyle's law:  $P_1 \times V_1 = P_2 \times V_2$
- Charles's law:  $\frac{V_1}{T_1} = \frac{V_2}{T_2}$
- Gay-Lussac's law:  $\frac{P_1}{T_1} = \frac{P_2}{T_2}$
- Combined gas law:  $\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$
- Ideal gas law:  $P \times V = n \times R \times T$  or  $PV = nRT$
- Dalton's law:  $P_{\text{total}} = P_1 + P_2 + P_3 + \dots$
- Graham's law:  $\frac{\text{Rate}_A}{\text{Rate}_B} = \sqrt{\frac{\text{molar mass}_B}{\text{molar mass}_A}}$

## Organizing Information

Use these terms to construct a concept map that organizes the major ideas of this chapter.



**Concept Map 14** Solve the Concept Map with the help of an interactive guided tutorial.

with **ChemASAP**

Boyle's law

Charles's law

partial pressure

combined gas law

diffusion

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**Judd E. Hollander**



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