

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES  
(CHEMISTRY)****TABLE 1      PHYSICAL CONSTANTS**

| NAME                            | SYMBOL     | VALUE                                     |
|---------------------------------|------------|---|
| Magnitude of charge on electron | $e$        | $1,6 \times 10^{-19} \text{ C}$           |
| Mass of an electron             | $m_e$      | $9,1 \times 10^{-31} \text{ kg}$          |
| Standard pressure               | $p^\theta$ | $1,01 \times 10^5 \text{ Pa}$             |
| Molar gas volume at STP         | $V_m$      | $22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$ |
| Standard temperature            | $T^\theta$ | $273 \text{ K (0 } ^\circ\text{C)}$       |
| Avogadro's constant             | $N_A$      | $6,02 \times 10^{23} \text{ mol}^{-1}$    |
| Faraday's constant              | $F$        | $96\,500 \text{ C} \cdot \text{mol}^{-1}$ |

**TABLE 2      CHEMISTRY FORMULAE**

|   |   |                     |
|---|---|---------------------|
| $n = \frac{m}{M}$                       | $n = \frac{N}{N_A}$   | $n = \frac{V}{V_m}$ |
| $c = \frac{n}{V}$ OR $c = \frac{m}{MV}$ | $K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$<br>at $25\, ^\circ\text{C (298 K)}$  |                     |
| $q = It$<br>$q = nF$                    | $E_{\text{cell}}^\theta = E_{\text{cathode}}^\theta - E_{\text{anode}}^\theta$<br>$E_{\text{cell}}^\theta = E_{\text{oxidising agent}}^\theta - E_{\text{reducing agent}}^\theta$ |                     |

# Physical Science P2 Nov2014

**Mary Somerville**



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