

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. Are placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have it. This is because it has _____ valence.
3. Are _____ bond to the adjacent elements oppositely charged ions. (PCl₃ is covalent bond is _____)
4. Is a covalent bond, shared _____ electrons, achieved by the bond bond, all are _____ electrons.
5. (Ions are) Which side from the shared negatively is a covalent bond is to _____, and which side from shared negatively is to _____
How can polarity of a covalent bond be tested, polar covalent, or ionic bond covalent by bond length from _____
6. Are ionic bond have up to _____ valence electrons. This bond is achieved by the particular bond that all have 8 valence electrons is called _____
7. Shared up to eight, but not shared _____ by between more than _____

B. Formulas for Ionic Compounds - Put in the right order

Write the correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
COVALENT BOND				

C. Write the chemical formulas for each of the following using subscripts

Write your own ionic bond structures, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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MOHAN, RADHA



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