Directions: Match each term in Column I with its description in Column II. Write the letter of the correct term in the space provided.

Column I binary compound chemically stable 3. nonpolar molecule _____ 4. ion ____ 5. ionic bond 6. polar molecule 7. oxidation number 8. chemical formula 9. covalent bond _____ 10. hydrate ----- 11. chemical bond _____ 12. polyatomic ion

Column II

a. number that indicates how many electrons an atom must gain, lose, or share to become stable

Class

- b. shorthand that tells what elements a compound contains and the exact number of atoms of each element in a unit of the compound
- c. positively or negatively charged, covalently bonded group of atoms
- d. compound composed of two elements
- e. describes an atom that has a full outermost energy level
- f. molecule that has a slightly positive end and a slightly negative end
- g. the attraction that forms between atoms when they share electrons
- h. the force that holds atoms together in a compound
- a compound that has water chemically attached to it
- j. the force of attraction between the opposite charges of the ions in an ionic compound
- k. molecule made of two identical atoms that share the electrons equally
- a charged particle that has either more or fewer electrons than protons

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S Ashworth

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