Gas Stoichiometry Worksheet

Name:

Solve all the following gas law problems. Show all work, answers are given at the end of the problem.

Molar Volume

1. Calculate the number of moles contained in \$50.mL of carbon dioxide at STP. (0.0246mol)

2. Calculate the mass of 1.50 L of CH4 at STP. (1.07g)

Calculate the volume in liters of 50.0 grams of nitrogen dioxide at STP. (24.3L)

Gas Stoichiometry

Given the following reaction: NH_aNO₂(s) → N₂(g) + 2H₂O (g) How many liters of nitrogen gas is produced if 50.0L of water is produced at STP? (25.0L)

Given the following reaction: CuO (s) + H₂(g) → Cu (s) + H₂O (g) If 250. L of hydrogen gas are used to reduce copper (II) oxide at STP, what mass of copper is to be expected? (709g)

 How many grams of solid lithium must be added to liquid water in order to obtain 15.0 L of hydrogen gas at STP? (9.29g)
 2 Li + 2H₂O → 2LiOH + H₂

Self Check Answers Gas Stoichiometry

Steven S. Zumdahl

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Basic Chemistry Steven S Zumdahl,1996 General Chemistry Workbook Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material Introductory Chemistry Steven S. Zumdahl,1993

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