

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

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Table of Contents Section 3 Chemistry Answers

1. Understanding the eBook Section 3 Chemistry Answers
 - The Rise of Digital Reading Section 3 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 3 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 3 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 3 Chemistry Answers

- Personalized Recommendations
- Section 3 Chemistry Answers User Reviews and Ratings
- Section 3 Chemistry Answers and Bestseller Lists
- 5. Accessing Section 3 Chemistry Answers Free and Paid eBooks
 - Section 3 Chemistry Answers Public Domain eBooks
 - Section 3 Chemistry Answers eBook Subscription Services
 - Section 3 Chemistry Answers Budget-Friendly Options
- 6. Navigating Section 3 Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 3 Chemistry Answers Compatibility with Devices
 - Section 3 Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 3 Chemistry Answers
 - Highlighting and Note-Taking Section 3 Chemistry Answers
 - Interactive Elements Section 3 Chemistry Answers
- 8. Staying Engaged with Section 3 Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 3 Chemistry Answers
- 9. Balancing eBooks and Physical Books Section 3 Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 3 Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 3 Chemistry Answers
 - Setting Reading Goals Section 3 Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 3 Chemistry Answers

- Fact-Checking eBook Content of Section 3 Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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