

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

losing

true

true

polyatomic

seven

true (released)

9. Calcium becomes a monatomic cation by gaining two electrons.

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

12. Monatomic ions consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

Section Reinforcement Metals Answers

A Loxley



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Table of Contents Section Reinforcement Metals Answers

1. Understanding the eBook Section Reinforcement Metals Answers
 - The Rise of Digital Reading Section Reinforcement Metals Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section Reinforcement Metals Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section Reinforcement Metals Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section Reinforcement Metals Answers

- Personalized Recommendations
- Section Reinforcement Metals Answers User Reviews and Ratings
- Section Reinforcement Metals Answers and Bestseller Lists
- 5. Accessing Section Reinforcement Metals Answers Free and Paid eBooks
 - Section Reinforcement Metals Answers Public Domain eBooks
 - Section Reinforcement Metals Answers eBook Subscription Services
 - Section Reinforcement Metals Answers Budget-Friendly Options
- 6. Navigating Section Reinforcement Metals Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section Reinforcement Metals Answers Compatibility with Devices
 - Section Reinforcement Metals Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section Reinforcement Metals Answers
 - Highlighting and Note-Taking Section Reinforcement Metals Answers
 - Interactive Elements Section Reinforcement Metals Answers
- 8. Staying Engaged with Section Reinforcement Metals Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section Reinforcement Metals Answers
- 9. Balancing eBooks and Physical Books Section Reinforcement Metals Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section Reinforcement Metals Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section Reinforcement Metals Answers
 - Setting Reading Goals Section Reinforcement Metals Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Reinforcement Metals Answers

- Fact-Checking eBook Content of Section Reinforcement Metals Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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