

Chapter 5: Chemical changes 1

Knowledge organiser

Reactions of metals

The reactivity of a metal is how chemically reactive it is. It is often related to water, some metals react with oxygen - these metals have high reactivity. Other metals will barely react with water or not, so aren't as reactive at all - these metals have low reactivity.

Reaction with water	Reaction with acid	Reactivity series	Displacement reactions
Reacts, gives off hydrogen gas	Reacts	most reactive potassium sodium calcium magnesium aluminium zinc iron lead copper silver gold	more reactive
Reacts, gives off hydrogen gas	Reacts, gives off hydrogen gas		
Reacts, only slowly	Reacts, only slowly		
Reacts, slowly with water only	Reacts, slowly with water only		
Does not react	Does not react		less reactive than the element above

Metal extraction

Some metals, like gold, are so unreactive that they are found as pure metals in the Earth's crust and can be mined.

Most metals will be compounds in metal ores and have to be extracted from the rock. This is a process called metal extraction in the metal to be found extracting it is called an ore.

Metals that are less reactive than carbon can be extracted by reduction with carbon. For example:

iron oxide + carbon $\rightarrow$ iron + carbon dioxide

Metals that are more reactive than carbon can be extracted using a process called electrolysis.

Salts

When acids react with metals or metal compounds, they form salts. It can be a compound where the hydrogen from an acid has been replaced by a metal. For example, when acid, HNO_3 , reacts with sodium hydroxide, NaOH , they will combine and be replaced with Na.

This table shows how to name salts.

Base	Acids (name acid)	Salts (name salt)	Other (name)
Hydroxide	OH^- and H^+	OH^- and Na^+	NaOH
Sulfate (from sulfuric acid)	SO_4^{2-} and H^+	SO_4^{2-} and Na^+	Na_2SO_4
Carbonate (from carbonic acid)	CO_3^{2-} and H^+	CO_3^{2-} and Na^+	Na_2CO_3
Chloride (from hydrochloric acid)	Cl^- and H^+	Cl^- and Na^+	NaCl

Chapter 5: Chemical changes 2

Knowledge organiser

Reactions of acids

Reactions of acids with metals

Acids react with some metals to form salts and hydrogen gas.

Reactions of acids with metal hydroxides

Acids react with metal hydroxides to form salts and water.

The ionic equation for this reaction is always:

$\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$

Reactions of acids with metal oxides

Acids react with metal oxides to form salts and water.

The ionic equation for this reaction is always:

$\text{H}^+ + \text{O}^{2-} \rightarrow \text{O}^{2-} + 2\text{H}^+ \rightarrow \text{H}_2\text{O}$

Balancing symbol equations

When writing symbol equations you need to ensure that the number of each atom is the same on both sides.

Example 1: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

There are 2 hydrogen atoms on each side, but 1 oxygen atom on the left and 2 on the right.

Example 2: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Example 3: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Example 4: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Example 5: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Example 6: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Example 7: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

There are 2 hydrogen atoms on each side, and 2 oxygen atoms on each side.

Alkalis and bases

Some reactions with acids form water in neutralisation reactions. Some metal hydroxides dissolve in water to form alkali solutions, called alkalis.

Some metal oxides and metal hydroxides do not dissolve in water. They are bases, but are not alkalis.

State symbols

A balanced symbol equation should also include state symbols.

Example 1: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

Example 2: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}_2$

Example 3: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 4: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 5: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

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Example 7: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 8: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 9: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 10: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 11: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 12: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Example 13: $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{H}_2\text{O}_2$

Displacement reactions

In a displacement reaction a more reactive element replaces the place of a less reactive element in a compound.

Example 1: $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$

Example 2: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Example 3: $\text{Al} + \text{CuSO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{Cu}$

Example 4: $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$

Example 5: $\text{Ca} + \text{CuSO}_4 \rightarrow \text{CaSO}_4 + \text{Cu}$

Example 6: $\text{Na} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu}$

Example 7: $\text{K} + \text{CuSO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{Cu}$

Example 8: $\text{Ba} + \text{CuSO}_4 \rightarrow \text{BaSO}_4 + \text{Cu}$

Example 9: $\text{Sr} + \text{CuSO}_4 \rightarrow \text{SrSO}_4 + \text{Cu}$

Example 10: $\text{Rb} + \text{CuSO}_4 \rightarrow \text{Rb}_2\text{SO}_4 + \text{Cu}$

Example 11: $\text{Cs} + \text{CuSO}_4 \rightarrow \text{Cs}_2\text{SO}_4 + \text{Cu}$

Example 12: $\text{Fr} + \text{CuSO}_4 \rightarrow \text{Fr}_2\text{SO}_4 + \text{Cu}$

Example 13: $\text{Ra} + \text{CuSO}_4 \rightarrow \text{RaSO}_4 + \text{Cu}$

Example 14: $\text{Ac} + \text{CuSO}_4 \rightarrow \text{Ac}_2\text{SO}_4 + \text{Cu}$

Example 15: $\text{Th} + \text{CuSO}_4 \rightarrow \text{ThSO}_4 + \text{Cu}$

Example 16: $\text{Pa} + \text{CuSO}_4 \rightarrow \text{Pa}_2\text{SO}_7 + \text{Cu}$

Example 17: $\text{U} + \text{CuSO}_4 \rightarrow \text{U}_3\text{O}_8 + \text{Cu}$

Example 18: $\text{Np} + \text{CuSO}_4 \rightarrow \text{Np}_2\text{O}_7 + \text{Cu}$

Example 19: $\text{Pu} + \text{CuSO}_4 \rightarrow \text{Pu}_2\text{O}_7 + \text{Cu}$

Example 20: $\text{Am} + \text{CuSO}_4 \rightarrow \text{Am}_2\text{O}_3 + \text{Cu}$

Example 21: $\text{Cm} + \text{CuSO}_4 \rightarrow \text{Cm}_2\text{O}_3 + \text{Cu}$

Example 22: $\text{Bk} + \text{CuSO}_4 \rightarrow \text{Bk}_2\text{O}_3 + \text{Cu}$

Example 23: $\text{Hf} + \text{CuSO}_4 \rightarrow \text{HfO}_2 + \text{Cu}$

Example 24: $\text{Ta} + \text{CuSO}_4 \rightarrow \text{Ta}_2\text{O}_5 + \text{Cu}$

Example 25: $\text{W} + \text{CuSO}_4 \rightarrow \text{WO}_3 + \text{Cu}$

Example 26: $\text{Re} + \text{CuSO}_4 \rightarrow \text{Re}_2\text{O}_7 + \text{Cu}$

Example 27: $\text{Os} + \text{CuSO}_4 \rightarrow \text{OsO}_4 + \text{Cu}$

Example 28: $\text{Ir} + \text{CuSO}_4 \rightarrow \text{Ir}_2\text{O}_3 + \text{Cu}$

Example 29: $\text{Pt} + \text{CuSO}_4 \rightarrow \text{PtO}_2 + \text{Cu}$

Example 30: $\text{Au} + \text{CuSO}_4 \rightarrow \text{Au}_2\text{O}_3 + \text{Cu}$

Example 31: $\text{Hg} + \text{CuSO}_4 \rightarrow \text{Hg}_2\text{SO}_4 + \text{Cu}$

Reactivity and ions

In a displacement reaction a more reactive element replaces the place of a less reactive element in a compound.

Example 1: $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$

Example 2: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Example 3: $\text{Al} + \text{CuSO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{Cu}$

Example 4: $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$

Example 5: $\text{Ca} + \text{CuSO}_4 \rightarrow \text{CaSO}_4 + \text{Cu}$

Example 6: $\text{Na} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu}$

Example 7: $\text{K} + \text{CuSO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{Cu}$

Example 8: $\text{Ba} + \text{CuSO}_4 \rightarrow \text{BaSO}_4 + \text{Cu}$

Example 9: $\text{Sr} + \text{CuSO}_4 \rightarrow \text{SrSO}_4 + \text{Cu}$

Example 10: $\text{Rb} + \text{CuSO}_4 \rightarrow \text{Rb}_2\text{SO}_4 + \text{Cu}$

Example 11: $\text{Cs} + \text{CuSO}_4 \rightarrow \text{Cs}_2\text{SO}_4 + \text{Cu}$

Example 12: $\text{Fr} + \text{CuSO}_4 \rightarrow \text{Fr}_2\text{SO}_4 + \text{Cu}$

Example 13: $\text{Ra} + \text{CuSO}_4 \rightarrow \text{RaSO}_4 + \text{Cu}$

Example 14: $\text{Ac} + \text{CuSO}_4 \rightarrow \text{Ac}_2\text{SO}_4 + \text{Cu}$

Example 15: $\text{Th} + \text{CuSO}_4 \rightarrow \text{ThSO}_4 + \text{Cu}$

Example 16: $\text{Pa} + \text{CuSO}_4 \rightarrow \text{Pa}_2\text{O}_7 + \text{Cu}$

Example 17: $\text{U} + \text{CuSO}_4 \rightarrow \text{U}_3\text{O}_8 + \text{Cu}$

Example 18: $\text{Np} + \text{CuSO}_4 \rightarrow \text{Np}_2\text{O}_7 + \text{Cu}$

Example 19: $\text{Pu} + \text{CuSO}_4 \rightarrow \text{Pu}_2\text{O}_7 + \text{Cu}$

Example 20: $\text{Am} + \text{CuSO}_4 \rightarrow \text{Am}_2\text{O}_3 + \text{Cu}$

Example 21: $\text{Cm} + \text{CuSO}_4 \rightarrow \text{Cm}_2\text{O}_3 + \text{Cu}$

Example 22: $\text{Bk} + \text{CuSO}_4 \rightarrow \text{Bk}_2\text{O}_3 + \text{Cu}$

Example 23: $\text{Hf} + \text{CuSO}_4 \rightarrow \text{HfO}_2 + \text{Cu}$

Example 24: $\text{Ta} + \text{CuSO}_4 \rightarrow \text{Ta}_2\text{O}_5 + \text{Cu}$

Example 25: $\text{W} + \text{CuSO}_4 \rightarrow \text{WO}_3 + \text{Cu}$

Example 26: $\text{Re} + \text{CuSO}_4 \rightarrow \text{Re}_2\text{O}_7 + \text{Cu}$

Example 27: $\text{Os} + \text{CuSO}_4 \rightarrow \text{OsO}_4 + \text{Cu}$

Example 28: $\text{Ir} + \text{CuSO}_4 \rightarrow \text{Ir}_2\text{O}_3 + \text{Cu}$

Example 29: $\text{Pt} + \text{CuSO}_4 \rightarrow \text{PtO}_2 + \text{Cu}$

Example 30: $\text{Au} + \text{CuSO}_4 \rightarrow \text{Au}_2\text{O}_3 + \text{Cu}$

Example 31: $\text{Hg} + \text{CuSO}_4 \rightarrow \text{Hg}_2\text{SO}_4 + \text{Cu}$

Half equations (HIT only)

In a displacement reaction a more reactive element replaces the place of a less reactive element in a compound.

Example 1: $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$

Example 2: $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$

Example 3: $\text{Al} + \text{CuSO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{Cu}$

Example 4: $\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$

Example 5: $\text{Ca} + \text{CuSO}_4 \rightarrow \text{CaSO}_4 + \text{Cu}$

Example 6: $\text{Na} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu}$

Example 7: $\text{K} + \text{CuSO}_4 \rightarrow \text{K}_2\text{SO}_4 + \text{Cu}$

Example 8: $\text{Ba} + \text{CuSO}_4 \rightarrow \text{BaSO}_4 + \text{Cu}$

Example 9: $\text{Sr} + \text{CuSO}_4 \rightarrow \text{SrSO}_4 + \text{Cu}$

Example 10: $\text{Rb} + \text{CuSO}_4 \rightarrow \text{Rb}_2\text{SO}_4 + \text{Cu}$

Example 11: $\text{Cs} + \text{CuSO}_4 \rightarrow \text{Cs}_2\text{SO}_4 + \text{Cu}$

Example 12: $\text{Fr} + \text{CuSO}_4 \rightarrow \text{Fr}_2\text{SO}_4 + \text{Cu}$

Example 13: $\text{Ra} + \text{CuSO}_4 \rightarrow \text{RaSO}_4 + \text{Cu}$

Example 14: $\text{Ac} + \text{CuSO}_4 \rightarrow \text{Ac}_2\text{SO}_4 + \text{Cu}$

Example 15: $\text{Th} + \text{CuSO}_4 \rightarrow \text{ThSO}_4 + \text{Cu}$

Example 16: $\text{Pa} + \text{CuSO}_4 \rightarrow \text{Pa}_2\text{O}_7 + \text{Cu}$

Example 17: $\text{U} + \text{CuSO}_4 \rightarrow \text{U}_3\text{O}_8 + \text{Cu}$

Example 18: $\text{Np} + \text{CuSO}_4 \rightarrow \text{Np}_2\text{O}_7 + \text{Cu}$

Example 19: $\text{Pu} + \text{CuSO}_4 \rightarrow \text{Pu}_2\text{O}_7 + \text{Cu}$

Example 20: $\text{Am} + \text{CuSO}_4 \rightarrow \text{Am}_2\text{O}_3 + \text{Cu}$

Example 21: $\text{Cm} + \text{CuSO}_4 \rightarrow \text{Cm}_2\text{O}_3 + \text{Cu}$

Example 22: $\text{Bk} + \text{CuSO}_4 \rightarrow \text{Bk}_2\text{O}_3 + \text{Cu}$

Example 23: $\text{Hf} + \text{CuSO}_4 \rightarrow \text{HfO}_2 + \text{Cu}$

Example 24: $\text{Ta} + \text{CuSO}_4 \rightarrow \text{Ta}_2\text{O}_5 + \text{Cu}$

Example 25: $\text{W} + \text{CuSO}_4 \rightarrow \text{WO}_3 + \text{Cu}$

Steps for writing an ionic equation (HIT only)

1. Write the full ionic equation.

2. Cancel out the same ions on both sides.

3. Write the net ionic equation.

4. Check the equation is balanced.

5. Add state symbols to the equation.

6. Add the net ionic equation to the full ionic equation.

7. Add the net ionic equation to the full ionic equation.

8. Add the net ionic equation to the full ionic equation.

9. Add the net ionic equation to the full ionic equation.

10. Add the net ionic equation to the full ionic equation.

11. Add the net ionic equation to the full ionic equation.

12. Add the net ionic equation to the full ionic equation.

13. Add the net ionic equation to the full ionic equation.

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15. Add the net ionic equation to the full ionic equation.

16. Add the net ionic equation to the full ionic equation.

17. Add the net ionic equation to the full ionic equation.

18. Add the

Section Chemical Changes Answers

Michael Seilmaier

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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