

7. molar heat of condensation

Part B True-False

8. AT 10. ST 12. NT
9. NT 11. NT

Part C Matching

13. a 15. b 17. d
14. e 16. c

Part D Questions and Problems

18. a. endothermic d. endothermic
b. endothermic e. exothermic
c. exothermic f. exothermic

$$19. \Delta H = 28.3 \frac{\text{g H}_2\text{O}(g)}{1 \text{ mol H}_2\text{O}(g)} \times \frac{1 \text{ mol H}_2\text{O}(g)}{18.0 \text{ g H}_2\text{O}(g)} \\ \times \frac{6.01 \text{ kJ}}{1 \text{ mol H}_2\text{O}(g)} = 9.45 \text{ kJ}$$

$$20. \Delta H = 5.53 \frac{\text{mol NH}_4\text{NO}_3(g)}{1 \text{ mol NH}_4\text{NO}_3(g)} \\ \times \frac{25.7 \text{ kJ}}{1 \text{ mol NH}_4\text{NO}_3(g)} = 142 \text{ kJ}$$

Section 17.4

Part A Completion

- sum
- enthalpy
- indirectly
- changed (reversed)
- standard heat of formation
- change
- one
- ΔH_f°
- zero
- subtracting

Part B True-False

11. NT 13. AT 15. NT
12. NT 14. NT

Part C Matching

16. b 18. a 20. c
17. c 19. d

Part D Questions and Problems



$$22. \Delta H_f^\circ (\text{products}) \\ = 3 \frac{\text{mol CO}_2(g)}{1 \text{ mol CO}_2(g)} \times \frac{-393.5 \text{ kJ}}{1 \text{ mol CO}_2(g)} \\ = -1181 \text{ kJ} \\ \Delta H_f^\circ (\text{reactants}) \\ = 3 \frac{\text{mol CO}(g)}{1 \text{ mol CO}(g)} \times \frac{-110.5 \text{ kJ}}{1 \text{ mol CO}(g)} + (-822.1 \text{ kJ}) \\ = -1154 \text{ kJ} \\ \Delta H^\circ = \Delta H_f^\circ (\text{products}) - \Delta H_f^\circ (\text{reactants}) \\ = -1181 \text{ kJ} - (-1154 \text{ kJ}) \\ = -27 \text{ kJ}$$

Practice Problems

Section 17.1

$$1. 200.0 \text{ Cal} \times \frac{1000 \text{ cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ cal}} \times \frac{1 \text{ kJ}}{10^3 \text{ J}} \\ = 836.8 \text{ kJ} \\ 2. C = \frac{525.0 \text{ cal}}{25.0 \text{ g} \times 15.0^\circ\text{C}} = 1.40 \text{ cal/(g} \times ^\circ\text{C)} \\ 3. \Delta T = \frac{1255.0 \text{ J}}{100.0 \text{ g} \times 2.1 \text{ J/(g} \times ^\circ\text{C)}} = 6.0^\circ\text{C} \\ 4. q = 100.0 \text{ g} \times 120.0^\circ\text{C} \times 0.90 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \\ = 1.1 \times 10^4 \text{ J}$$

Section 17.2

$$1. \Delta H = 150.0 \text{ g} \times 4.18 \frac{\text{J}}{\text{g} \times ^\circ\text{C}} \times 10^\circ\text{C} \\ = 6.3 \times 10^3 \text{ J} = 6.3 \text{ kJ} \\ 2. \Delta H = 15.0 \text{ g Ca(OH)}_2(g) \times \frac{1 \text{ mol Ca(OH)}_2(g)}{74.1 \text{ g Ca(OH)}_2(g)} \\ \times \frac{-65.2 \text{ kJ}}{1 \text{ mol Ca(OH)}_2(g)} = -13.2 \text{ kJ} \\ 3. \Delta H = 52.4 \text{ g CH}_4(g) \times \frac{1 \text{ mol CH}_4(g)}{16.0 \text{ g CH}_4(g)} \\ \times \frac{-890.2 \text{ kJ}}{1 \text{ mol CH}_4(g)} = 2.93 \times 10^3 \text{ kJ} \\ 4. 4\text{NH}_3 + 5\text{O}_2(g) \rightarrow 4\text{NO}(g) + 6\text{H}_2\text{O}(g) \\ \Delta H = \frac{-226 \text{ kJ}}{1 \text{ mol NH}_3(g)} \times 4 \text{ mol NH}_3(g) \\ = -904 \text{ kJ}$$

Prentice Hall Chemistry Answers Chapter 2

Justin P. Lomont, Ian C. Stewart



Prentice Hall Chemistry Answers Chapter 2:

Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science, 2003-11
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Prentice Hall Biology, 2002 Irvine Welsh, Prentice-Hall Staff, 2001-04 One program that ensures success for all students

Electrochemistry Christine Lefrou, Pierre Fabry, Jean-Claude Poignet, 2012-05-24 This textbook offers original and new approaches to the teaching of electrochemical concepts principles and applications Throughout the text the authors provide a balanced coverage of the thermodynamic and kinetic processes at the heart of electrochemical systems The first half of the book outlines fundamental concepts appropriate to undergraduate students and the second half gives an in depth account of electrochemical systems suitable for experienced scientists and course lecturers Concepts are clearly explained and mathematical treatments are kept to a minimum or reported in appendices This book features Questions and answers for self assessment Basic and advanced level numerical descriptions Illustrated electrochemistry applications This book is accessible to both novice and experienced electrochemists and supports a deep understanding of the fundamental principles and laws of electrochemistry

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Table of Contents Prentice Hall Chemistry Answers Chapter 2

1. Understanding the eBook Prentice Hall Chemistry Answers Chapter 2
 - The Rise of Digital Reading Prentice Hall Chemistry Answers Chapter 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry Answers Chapter 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry Answers Chapter 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry Answers Chapter 2
 - Personalized Recommendations
 - Prentice Hall Chemistry Answers Chapter 2 User Reviews and Ratings
 - Prentice Hall Chemistry Answers Chapter 2 and Bestseller Lists

5. Accessing Prentice Hall Chemistry Answers Chapter 2 Free and Paid eBooks
 - Prentice Hall Chemistry Answers Chapter 2 Public Domain eBooks
 - Prentice Hall Chemistry Answers Chapter 2 eBook Subscription Services
 - Prentice Hall Chemistry Answers Chapter 2 Budget-Friendly Options
6. Navigating Prentice Hall Chemistry Answers Chapter 2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry Answers Chapter 2 Compatibility with Devices
 - Prentice Hall Chemistry Answers Chapter 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry Answers Chapter 2
 - Highlighting and Note-Taking Prentice Hall Chemistry Answers Chapter 2
 - Interactive Elements Prentice Hall Chemistry Answers Chapter 2
8. Staying Engaged with Prentice Hall Chemistry Answers Chapter 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry Answers Chapter 2
9. Balancing eBooks and Physical Books Prentice Hall Chemistry Answers Chapter 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry Answers Chapter 2
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry Answers Chapter 2
 - Setting Reading Goals Prentice Hall Chemistry Answers Chapter 2
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry Answers Chapter 2
 - Fact-Checking eBook Content of Prentice Hall Chemistry Answers Chapter 2
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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