

When $k_1 = k_2$ the (Counterintuitive?) result is that the same size of MFR and PFR will maximize the concentration of R (the intermediate product): $C_{R, \max}$ is higher in the PFR, but the volumes are the same!

My solution (partly graphical) showed that $V_{PFR} \approx 153 \text{ L}$ while $V_{MFR} \approx 167 \text{ L}$. In fact, they should be identical! ^(?)

Here is the justification (after Kelly Caputo pointed it out to me... Thanks, Kelly!!!):

$$\tau_{MFR, \text{opt}} = \frac{1}{\sqrt{R_1 R_2}} = \underline{10 \text{ min (OK)}} \quad \frac{V_{PFR} = V_{MFR} = 167 \text{ L}}{\nearrow}$$

(eq. 8.15)

For the PFR, we use L'Hôpital's theorem

$$\tau_{PFR, \text{opt}} = \lim_{R_2/k_1 \rightarrow 1} \frac{\ln\left(\frac{R_2}{R_1}\right)}{R_2 - R_1} = \frac{\lim\left(1 + \ln\frac{R_2}{R_1}\right)}{\lim(R_2)} = \frac{1}{R_2} = \underline{10 \text{ min}}$$

So, the important comment here, in agreement with Levenspiel's statement on p.176, is NOT ONLY that the MFR is "not much worse than" the PFR, But that in terms of activity - even though conversion is lower - the MFR does not have the typical size disadvantage as well. (vs. selectivity)

Please add this page to your [HW #6] Solution package!

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web core concepts some reactions will go fast and some will go slow the speed of the reaction is it s reaction rate which is dictated by a rate law in this article we will learn about reaction rates rate laws the rate constant and the reaction order

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the first is a problem of applied thermodynamics the second is the central problem of chemical kinetics

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we can distinguish two general approaches to this task one approach is to measure reaction rate directly that is for a b c a b
c we measure the reaction rate in experiments where the concentrations a a

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to be text rate k left ce CH_3OH right nonumber what is the order of reaction with respect to methanol and ethyl acetate and
what is the overall order of reaction

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equation for a chemical reaction is an expression that provides a relationship between the rate of the reaction and the
concentrations of the reactants participating in it table of contents expression reaction